

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction:

This experiment involves the gravimetric analysis of a sample solution containing chloride ions. The sample solution is a mixture of sodium chloride and other soluble salts. The sample is weighed and then dissolved in water. The solution is then treated with silver nitrate, which reacts with the chloride ions to form a precipitate of silver chloride. The precipitate is then filtered, washed, and dried. The mass of the precipitate is then measured, and the mass of chloride in the original sample is calculated. The experiment is performed in a fume hood to avoid exposure to silver nitrate, which is a strong oxidizing agent and can cause skin irritation. The experiment is also performed in a clean, dry environment to avoid contamination of the sample.

The sample is weighed and then dissolved in water. The solution is then treated with silver nitrate, which reacts with the chloride ions to form a precipitate of silver chloride. The precipitate is then filtered, washed, and dried. The mass of the precipitate is then measured, and the mass of chloride in the original sample is calculated.

Procedure:

The procedure involves the following steps: 1. Weighing the sample. 2. Dissolving the sample in water. 3. Adding silver nitrate to the solution. 4. Filtering the precipitate. 5. Washing the precipitate. 6. Drying the precipitate. 7. Weighing the dried precipitate. 8. Calculating the mass of chloride in the sample.

Data Analysis:

	Sample 1	Sample 2
Mass of sample	0.500 g	0.500 g
Mass of precipitate	0.850 g	0.850 g
Mass of chloride in sample	0.350 g	0.350 g

The mass of chloride in the sample is calculated as follows: $m_{\text{Cl}^-} = m_{\text{precipitate}} \times \frac{m_{\text{Cl}^-}}{m_{\text{AgCl}}}$

For Sample 1: $m_{\text{Cl}^-} = 0.850 \text{ g} \times \frac{35.45 \text{ g/mol}}{143.32 \text{ g/mol}} = 0.350 \text{ g}$

For Sample 2: $m_{\text{Cl}^-} = 0.850 \text{ g} \times \frac{35.45 \text{ g/mol}}{143.32 \text{ g/mol}} = 0.350 \text{ g}$

Analysis Of Chloride Experiment 7 Lab Report

L Cohen



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Table of Contents Analysis Of Chloride Experiment 7 Lab Report

1. Understanding the eBook Analysis Of Chloride Experiment 7 Lab Report
 - The Rise of Digital Reading Analysis Of Chloride Experiment 7 Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Analysis Of Chloride Experiment 7 Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Analysis Of Chloride Experiment 7 Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Analysis Of Chloride Experiment 7 Lab Report
 - Personalized Recommendations
 - Analysis Of Chloride Experiment 7 Lab Report User Reviews and Ratings
 - Analysis Of Chloride Experiment 7 Lab Report and Bestseller Lists
5. Accessing Analysis Of Chloride Experiment 7 Lab Report Free and Paid eBooks

- Analysis Of Chloride Experiment 7 Lab Report Public Domain eBooks
- Analysis Of Chloride Experiment 7 Lab Report eBook Subscription Services
- Analysis Of Chloride Experiment 7 Lab Report Budget-Friendly Options
- 6. Navigating Analysis Of Chloride Experiment 7 Lab Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Analysis Of Chloride Experiment 7 Lab Report Compatibility with Devices
 - Analysis Of Chloride Experiment 7 Lab Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Analysis Of Chloride Experiment 7 Lab Report
 - Highlighting and Note-Taking Analysis Of Chloride Experiment 7 Lab Report
 - Interactive Elements Analysis Of Chloride Experiment 7 Lab Report
- 8. Staying Engaged with Analysis Of Chloride Experiment 7 Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Analysis Of Chloride Experiment 7 Lab Report
- 9. Balancing eBooks and Physical Books Analysis Of Chloride Experiment 7 Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Analysis Of Chloride Experiment 7 Lab Report
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Analysis Of Chloride Experiment 7 Lab Report
 - Setting Reading Goals Analysis Of Chloride Experiment 7 Lab Report
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Analysis Of Chloride Experiment 7 Lab Report
 - Fact-Checking eBook Content of Analysis Of Chloride Experiment 7 Lab Report
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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