Reactivity of Alkaline Earth Metals and Alkali Metals Lab

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The elements in group 2 of the Periodic Table are called the Alkaline Earth metals. These were called the earths by early chemists because they were first isolated from compounds found in the earth. The metals were combined with oxygen. The alkaline part of the name comes from the fact that they firm have plus or abuse solutions in water. The group is composed of berylines, magnetium, calcium, threatings. Jacobs and safetime.

All alkaline earth metals have It valence electrons, which they tend to give up early, making them quite reactive. They are in reactive that they are never found uncombined in nature. In earlier for these ships white metals to remain in their superacted state, they must be pretected from all and water. May remain and calcium, for example, are obtained in their elemental visits by a channel proper railed electrology and then there is no protected in their elemental visits by a channel proper railed electrology and then thereof in anothers.

In this investigation you will explore the reactivity of magnessium, and calcium, two of the more common abalities earth metals. You will then compare these metals, with alternature, a member of group 2.

The elements in proup 1 of the Periodic Table are called the albah metals. These metals are even more reactive than the Albahus Earth metals and to they are also not found in their elemental form. The proup is composed of lithium, softum, potavitum, religious, and decision.

All Alkals metals have I valence electron, which they give up to readily that these demonstrations cannot be performed without significant safety preparations.

Printed Males and

How do the activities of the alliabos earth metals magnesium and calcium compare to the activity of alternation?

Market States

Copyles Marrepipettes Waste backer Ca metal furness flumer 5 plans Petri Cales Porceps steel wool 5 triber paper towels 10g ribbon.

Town Phenology Daleis Al Rul

Presidente.

PARTA

- 1. Part on Suffetty Clauses at
- Wash the Petrs dishes with water and dry completely with paper towel. Femore any white residue from the dishes.
- Examine the magnetium ribbon. If it is dull in order, solich it using steel worl.
- Using a permanent market, write "Mg" on one Petrs Each COVER, "Ca" on a second Petrs Each. COVER and "All on the last Petrs Each COVER.
- Safety: Any skin or elething that comes in contact with Ca should be washed.
 Instedicts by with large amounts of water. Make more there are no open flames when most limb Co with motor becomes the one can be seen feature.
- resulting Ca with water because the gas can be explosive.

 6. Open the Petri delice and add a packs of hig ribbon, Ca turning and Alfod to REPARATE CONTROL of each correctly labeled disk. Wrote descriptions of these metals in data table 1. Securing on differences between the tipes metals.
- 7. Add 2 pigestion full of water to each <u>Enthron</u> pater dark. Use the cover of each Peter dark to drop the metal nample into the water. COVER THE DISH EMMEDIATELY WITH ITS COVER, for not allow the water to touch your abin during or after the marking.
- W. Franced descriptions of any reactions in data table 1. If any metal is unreacted, sidd 1 more piperts full of water to make ourse it doesn't reach. When all reactions are done, open the Petri dishes and add 1 drop of phenolphthalasis, onto each piece of metal. Record observations of color changes in data table 1. Use the foreign and remove the Mg and Allfron the Petri dish and day them. with a paper toward to reach a part II.
- Dump the contents of the "Ca" Petri Ink into the <u>wants healest</u> at the teacher lab table. Fires out all THERES of the Petri Inkes with value. Doe Petri Salves with panel toward.

Observations of Alicaline Earth Metals:

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