

3 This question is about atoms, molecules and nanoparticles.

3 (a) Different atoms have different numbers of sub-atomic particles.

3 (a) (i) An oxygen atom can be represented as $^{16}_8\text{O}$

Explain why the mass number of this atom is 16.

You should refer to the numbers of sub-atomic particles in the nucleus of the atom.

[2 marks]

3 (a) (ii) Explain why $^{12}_6\text{C}$ and $^{14}_6\text{C}$ are isotopes of carbon.

You should refer to the numbers of sub-atomic particles in the nucleus of each isotope.

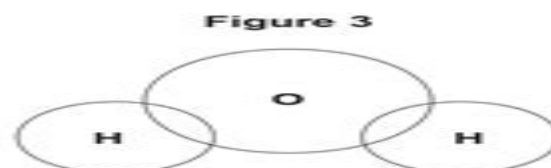
[3 marks]

3 (b) Hydrogen atoms and oxygen atoms chemically combine to produce water molecules.

3 (b) (i) Complete Figure 3 to show the arrangement of the outer shell electrons of the hydrogen and oxygen atoms in a molecule of water.

Use dots (•) or crosses (×) to represent the electrons.

[2 marks]



3 (b) (ii) Name the type of bonding in a molecule of water.

[1 mark]

3 (b) (iii) Why does pure water **not** conduct electricity?

[1 mark]

3 (c) Nanoparticles of cobalt oxide can be used as catalysts in the production of hydrogen from water.

3 (c) (i) How does the size of a nanoparticle compare with the size of an atom?

[1 mark]

3 (c) (ii) Suggest **one** reason why 1 g of cobalt oxide nanoparticles is a better catalyst than 1 g of cobalt oxide powder.

[1 mark]

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Yi-Tong Ma



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Table of Contents Alternative To Practical Chemistry Past Question 2013 Gsce

1. Understanding the eBook Alternative To Practical Chemistry Past Question 2013 Gsce
 - The Rise of Digital Reading Alternative To Practical Chemistry Past Question 2013 Gsce
 - Advantages of eBooks Over Traditional Books
2. Identifying Alternative To Practical Chemistry Past Question 2013 Gsce
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Alternative To Practical Chemistry Past Question 2013 Gsce
 - User-Friendly Interface
4. Exploring eBook Recommendations from Alternative To Practical Chemistry Past Question 2013 Gsce

- Personalized Recommendations
- Alternative To Practical Chemistry Past Question 2013 Gsce User Reviews and Ratings
- Alternative To Practical Chemistry Past Question 2013 Gsce and Bestseller Lists
- 5. Accessing Alternative To Practical Chemistry Past Question 2013 Gsce Free and Paid eBooks
 - Alternative To Practical Chemistry Past Question 2013 Gsce Public Domain eBooks
 - Alternative To Practical Chemistry Past Question 2013 Gsce eBook Subscription Services
 - Alternative To Practical Chemistry Past Question 2013 Gsce Budget-Friendly Options
- 6. Navigating Alternative To Practical Chemistry Past Question 2013 Gsce eBook Formats
 - ePub, PDF, MOBI, and More
 - Alternative To Practical Chemistry Past Question 2013 Gsce Compatibility with Devices
 - Alternative To Practical Chemistry Past Question 2013 Gsce Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Alternative To Practical Chemistry Past Question 2013 Gsce
 - Highlighting and Note-Taking Alternative To Practical Chemistry Past Question 2013 Gsce
 - Interactive Elements Alternative To Practical Chemistry Past Question 2013 Gsce
- 8. Staying Engaged with Alternative To Practical Chemistry Past Question 2013 Gsce
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Alternative To Practical Chemistry Past Question 2013 Gsce
- 9. Balancing eBooks and Physical Books Alternative To Practical Chemistry Past Question 2013 Gsce
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Alternative To Practical Chemistry Past Question 2013 Gsce
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Alternative To Practical Chemistry Past Question 2013 Gsce
 - Setting Reading Goals Alternative To Practical Chemistry Past Question 2013 Gsce
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Alternative To Practical Chemistry Past Question 2013 Gsce

- Fact-Checking eBook Content of Alternative To Practical Chemistry Past Question 2013 Gsce
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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