

Determination of the Thickness of Aluminum Foil

Name _____

Chemistry, Period x _____

Date Exp performed _____

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Introduction

Aluminum foil is a common household material that is so thin it is impossible to measure directly using a ruler. It is however possible to determine the thickness of aluminum foil through the use of the density of aluminum. Density is mass divided by volume and the density of pure aluminum is well known to be 2.70 g/cm^3 . If the mass of a sample of aluminum foil is known, then it is possible to calculate the volume of the foil. In general volume is the product of length times width times height of an object. If we know the length and width of a piece of foil, we can then determine its height, or thickness.

Aluminum foil is pure elemental aluminum and consists of nothing but aluminum atoms lined up in a three dimensional grid work. For this lab, we will assume the aluminum atoms are stacked side by side and direction on top of each other. Using this assumption, we can determine the number of Aluminum atoms thick the foil is, as well as the number of aluminum atoms preset in a given piece of foil.

Because aluminum foil is so thin and atoms are so small, many of the quantities used in this experiment will be very small and best expressed using scientific notation. Scientific notation expresses all of the significant digits for a quantity as a number between 1 and 9.999... followed by a multiple of 10 to express the order of magnitude of the number. Positive exponents of 10 indicate very large numbers while negative exponents of 10 indicate very small numbers.

Purpose

- To determine the thickness of a piece of aluminum foil in centimeters.
- To determine the number of atoms thick the foil is.
- To determine the total number of aluminum atoms present in a piece of aluminum foil.

Hypothesis

The thickness of foil will probably be about a micrometer thick so for centimeters that means a thickness of $1 \times 10^{-4} \text{ cm}$. Given that the size of an atom is about 10^{-10} m , the number of atoms in a $1 \times 10^{-4} \text{ cm}$ will be 1×10^6 . For a 100 cm^2 piece of foil (10 cm by 10 cm), the number of aluminum atoms will be about 1×10^{12} .

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