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OWNER'S MANUAL

Japanese domestic models

3 Mitsubishi Colt Owners Manual

L Reisser



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although the c-cl bonds are rather polar the individual bond dipoles cancel one another in this symmetrical structure and does not have a net dipole moment

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 determine based on electronegativity and symmetry if the following molecules are polar or nonpolar
 hcn h₂c=cf₃ c₂f₄ c₂f₆ c₂h₂cl₂ c₂h₂br₂ c₂h₂f₂ n₂f₄ n₂f₂ q3

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 a. cl₂ b. clo₂ c. tecl₂ d. pcl₃ e. sef₄ f. ph₂ g. xef₂ answer problem 6.2.3.6.2.3

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web the relative electronegativities of the bonded atoms is h < c < cl and so the bond moments all point toward the cl end of the

molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as BF_3 trigonal planar CH_4 tetrahedral PF_5 trigonal bipyramidal and SF_6

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