

Stoichiometry Limiting Reactant

Answers

- ① How many grams of calcium phosphate can be produced from the reaction between 2.50 L of 0.250 M calcium chloride with an excess of phosphoric acid?



$$2.50 \text{ L CaCl}_2 \times \frac{0.250 \text{ mol CaCl}_2}{1 \text{ L CaCl}_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{3 \text{ mol CaCl}_2} \times \frac{310 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 64.6 \text{ g Ca}_3(\text{PO}_4)_2$$

- ② How many milliliters of 1.50 M nitric acid is required to react with 100.0 g of cuprous oxide?



$$100 \text{ g Cu}_2\text{O} \times \frac{1 \text{ mol Cu}_2\text{O}}{143 \text{ g Cu}_2\text{O}} \times \frac{14 \text{ mol HNO}_3}{3 \text{ mol Cu}_2\text{O}} \times \frac{1000 \text{ ml HNO}_3}{1.5 \text{ mol HNO}_3} = 2180 \text{ ml HNO}_3$$

- ③ Consider the following unbalanced reaction:



Balance the reaction. 3.25 g of NH_3 is allowed to react with 3.5 g of O_2 .

- a) Which reactant is the limiting reagent?

$$3.25 \text{ g NH}_3 \times \frac{1 \text{ mol NH}_3}{17.03 \text{ g NH}_3} \times \frac{4 \text{ mol NO}}{4 \text{ mol NH}_3} \times \frac{30.01 \text{ g NO}}{1 \text{ mol NO}} = 5.73 \text{ g NO}$$

$$3.5 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol NO}}{5 \text{ mol O}_2} \times \frac{30.01 \text{ g NO}}{1 \text{ mol NO}} = 2.63 \text{ g NO}$$

O_2 is the limiting reagent.

- b) How many grams of NO are formed?

2.63 g of NO are formed.

- c) How much excess reagent remains after the reaction?

$$3.5 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol NH}_3}{5 \text{ mol O}_2} \times \frac{17.03 \text{ g NH}_3}{1 \text{ mol NH}_3} = 1.49 \text{ g NH}_3 \text{ used}$$

$$3.25 \text{ g NH}_3 - 1.49 \text{ g NH}_3 = 1.76 \text{ g NH}_3 \text{ remains.}$$

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 - The Rise of Digital Reading Answer Sheet To The Stoichiometry Limiting Reagent
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Sheet To The Stoichiometry Limiting Reagent
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Sheet To The Stoichiometry Limiting Reagent
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Sheet To The Stoichiometry Limiting Reagent

- Personalized Recommendations
 - Answer Sheet To The Stoichiometry Limiting Reagent User Reviews and Ratings
 - Answer Sheet To The Stoichiometry Limiting Reagent and Bestseller Lists
5. Accessing Answer Sheet To The Stoichiometry Limiting Reagent Free and Paid eBooks
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 - Answer Sheet To The Stoichiometry Limiting Reagent Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Answer Sheet To The Stoichiometry Limiting Reagent
 - Highlighting and Note-Taking Answer Sheet To The Stoichiometry Limiting Reagent
 - Interactive Elements Answer Sheet To The Stoichiometry Limiting Reagent
 8. Staying Engaged with Answer Sheet To The Stoichiometry Limiting Reagent
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Sheet To The Stoichiometry Limiting Reagent
 9. Balancing eBooks and Physical Books Answer Sheet To The Stoichiometry Limiting Reagent
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Sheet To The Stoichiometry Limiting Reagent
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Answer Sheet To The Stoichiometry Limiting Reagent
 - Setting Reading Goals Answer Sheet To The Stoichiometry Limiting Reagent
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Answer Sheet To The Stoichiometry Limiting Reagent

- Fact-Checking eBook Content of Answer Sheet To The Stoichiometry Limiting Reagent
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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