

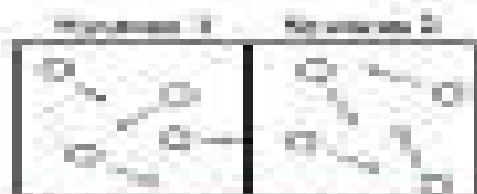
**Unit 6: Thermochemistry****6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



**Temperature:** is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with  $v$  is  $\frac{1}{2}mv^2$ , so the more of particles that move fast enough, the more kinetic energy we will see for a system. Therefore, the more particles move faster, the more kinetic energy we will see for the system. It's important to note that the more particles move faster, the more kinetic energy we have.

3. Sketch what you expect to represent the relative distribution of the particles in the following two parts: samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.  
System 1 is at a higher temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are moving faster, so the average KE is higher. In System 2, the particles are moving slower, so the average KE is lower. Therefore, System 1 is at a higher temperature than System 2.
- b. In which direction will kinetic energy be transferred? Justify your answer in terms of molecular collisions.  
Kinetic energy will be transferred from System 1 to System 2. In System 1, the particles are moving faster, so they have more kinetic energy. In System 2, the particles are moving slower, so they have less kinetic energy. When the two systems are in contact, the faster particles in System 1 will collide with the slower particles in System 2, transferring some of their kinetic energy to them. This will continue until the two systems reach thermal equilibrium.

# Answer Key Unit 6 Thermochemistry

**Darrell D. Ebbing**



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