

8. Sulfur, having atomic no. 16, is found within the Earth's crust. Sulfur is composed out of isotopes of almost all existing isotopes.

18 isotopes of sulfur from 18 isotopes were employed to give the following composition of isotopes.

isotopes	abundance (%)
^{32}S	95.02
^{34}S	4.29
^{36}S	0.69

- (a) Define the term relative atomic mass.

The relative atomic mass of an element is the average mass of its atoms, measured on the scale of ^{12}C isotope of carbon is 12.

- (b) Calculate the relative atomic mass of the isotopes of sulfur.

They give equation to find relative mass.

$$\left[\frac{32 \times 95.02}{100} \right] + \left[\frac{34 \times 4.29}{100} \right] + \left[\frac{36 \times 0.69}{100} \right]$$

$$= 32.07$$

$$\text{answer} = 32.07 \text{ amu}$$

- (c) Some isotopes are heavy from existing elements, however, they sometimes were made up of two particles which were called subatomic particles, electrons, protons and neutrons.

Complete the table to show the number of subatomic particles in the ^{32}S atom and $^{32}\text{S}^{2-}$ ion.

	protons	neutrons	electrons
^{32}S	16	16	16
$^{32}\text{S}^{2-}$	16	16	18



Answer To F321 Chemistry May Series

John Anderson



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