

Student Exploration: Calorimetry Lab

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance's **specific heat capacity**. You will use the Calorimetry Lab Gizmo™ to determine the specific heat capacities of various substances.



- On the **SIMULATION** pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the **GRAPH** tab, and click **Play** (▶).

A. What was the **Final temperature** of the copper and the water?

34.96°C

B. How much did the temperature of the copper change?

$$\begin{aligned} T &= T_f - T_i \\ &= (34.96^\circ\text{C} - 90^\circ\text{C}) \\ &= -55.04^\circ\text{C} \end{aligned}$$

The temperature of the copper changed about 55°C.

C. How much did the temperature of the water change?

$$\begin{aligned} T &= T_f - T_i \\ &= (34.96^\circ\text{C} - 30^\circ\text{C}) \\ &= 4.96^\circ\text{C} \end{aligned}$$

The temperature of the water changed about 5°C.

- Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

$$\begin{aligned} c_w &= 4.184 \text{ J/g} \\ c_c &= 0.385 \text{ J/g} \end{aligned}$$

Water has a greater specific heat capacity.

Activity A:

Heat transfer

Get the Gizmo ready:

- Click **Reset** (↺).



Question: What factors determine how heat energy transfers between objects?

- Predict: In the Gizmo warm-up, you saw how 200 g of 90 °C copper transfers heat to 200 g of 30.0 °C water.

A. How do you think increasing the water's mass would affect the final temperature?

Answers For Calorimetry Gizmo

Joint Committee on Calorimetry

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the text 'Joint Committee on Calorimetry'.

Answers For Calorimetry Gizmo:

The book delves into Answers For Calorimetry Gizmo. Answers For Calorimetry Gizmo is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Answers For Calorimetry Gizmo, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Answers For Calorimetry Gizmo
- Chapter 2: Essential Elements of Answers For Calorimetry Gizmo
- Chapter 3: Answers For Calorimetry Gizmo in Everyday Life
- Chapter 4: Answers For Calorimetry Gizmo in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Answers For Calorimetry Gizmo. The first chapter will explore what Answers For Calorimetry Gizmo is, why Answers For Calorimetry Gizmo is vital, and how to effectively learn about Answers For Calorimetry Gizmo.
3. In chapter 2, this book will delve into the foundational concepts of Answers For Calorimetry Gizmo. The second chapter will elucidate the essential principles that must be understood to grasp Answers For Calorimetry Gizmo in its entirety.
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5. In chapter 4, the author will scrutinize the relevance of Answers For Calorimetry Gizmo in specific contexts. The fourth chapter will explore how Answers For Calorimetry Gizmo is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Answers For Calorimetry Gizmo. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Answers For Calorimetry Gizmo.

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