lonic Bonding

Name: Hour:

Information: Polyatomic Ions

The word, "polyatomic" means "many atoms". A polyatomic ion, therefore, is an ion that is made of more than one atom. An example of a polyatomic ion is the sulfate ion, $SO_4{}^5$. Sulfate is composed of one sulfur atom and four oxygen atoms and overall sulfate has a negative two charge.

Some polyatomic ions:

Sulfate: SO₄²⁻ Phosphate: PO₄²⁻ Nitrate: NO₅²⁻ Cyanide: CN²⁻ Ammonium: NH₄²⁻ Chlorate: ClO₅²⁻ Acetate: C₂H₃O₂²⁻ Hydroxide: OH²⁻ Carbonate: CO₃²⁻

Critical Thinking Questions

- 1. What do all of the polyatomic ions that have the suffix "-ate" have in common?
- 2. Which two atoms do you think compose the polyatomic ion called "silicate"?
- 3. What is the difference between calcium nitride and calcium nitrate?

Information: Writing Formulas With Polyatomic Ions

First of all, you must remember that you can never change the formula for a polyatomic ion. Sulfate is <u>always</u> SO_4 and never S_2O_8 or something else. Following are some examples of chemical formulas that contain polyatomic ions.

Ammonium chloride is formed from one ammonium ion (NH₄*) and one chloride ion (Cl') to give the formula: NH₄Cl. Sodium sulfate requires two sodium ions (Na*) because sulfate (SO₄*) has a negative two charge; the formula is: Na₂SO₄.

Consider calcium hydroxide. Calcium has a positive two charge (Ca²⁺) and hydroxide has a negative one charge (OH'). We need two hydroxide ions to combine with one calcium ion so that the overall charge ends up being zero. We write calcium hydroxide like Ca(OH)₂.

Following are some more examples: potassium acetate: KC₂H₃O₂ magnesium nitrate: Mg(NO₃)₂ barium phosphate: Ba₃(PO₄)₂ calcium carbonate: CaCO₃

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Table of Contents Answers For Chemquest 20

- 1. Understanding the eBook Answers For Chemquest 20
 - The Rise of Digital Reading Answers For Chemquest 20
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Answers For Chemquest 20
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Chemquest 20
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Answers For Chemquest 20
 - Personalized Recommendations
 - Answers For Chemquest 20 User Reviews and Ratings

- Answers For Chemquest 20 and Bestseller Lists
- 5. Accessing Answers For Chemquest 20 Free and Paid eBooks
 - Answers For Chemquest 20 Public Domain eBooks
 - Answers For Chemquest 20 eBook Subscription Services
 - Answers For Chemquest 20 Budget-Friendly Options
- 6. Navigating Answers For Chemquest 20 eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Answers For Chemquest 20 Compatibility with Devices
 - Answers For Chemquest 20 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Chemquest 20
 - Highlighting and Note-Taking Answers For Chemquest 20
 - Interactive Elements Answers For Chemquest 20
- 8. Staying Engaged with Answers For Chemquest 20
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Chemquest 20
- 9. Balancing eBooks and Physical Books Answers For Chemquest 20
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Chemquest 20
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For Chemquest 20
 - Setting Reading Goals Answers For Chemquest 20
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers For Chemquest 20
 - Fact-Checking eBook Content of Answers For Chemquest 20
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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