

SCHAU Unit 4 - Quiz – Equilibrium

Name: _____ Period: _____

1. (20)

2. (20)

3. (20)

4. (20)

5. (20)

6. (20) (20) (20)

$$K_c = \frac{[\text{products}]}{[\text{reactants}]} \quad K_p = \frac{(P)^{\text{products}}}{(P)^{\text{reactants}}}$$

$$K_c = \frac{P_{\text{H}_2\text{O}}^2 P_{\text{CO}_2}}{P_{\text{H}_2}^2 P_{\text{CO}}} \quad K_p = \frac{(P_{\text{H}_2\text{O}})^2 (P_{\text{CO}_2})}{(P_{\text{H}_2})^2 (P_{\text{CO}})}$$

gases: $\Delta G^\circ = -RT \ln K$ $\Delta G^\circ = -RT \ln K_p$ $\Delta G^\circ = -RT \ln K_c$ $\Delta G^\circ = -RT \ln K$ $\Delta G^\circ = -RT \ln K$

Question 1

1. Hydrogen sulfide, H_2S , is a slightly poisonous gas. At 1000°C, an equilibrium mixture contains 0.020 moles of hydrogen, H_2 , 0.010 moles of hydrogen sulfide, H_2S , and an undetermined amount of sulfur, S_2 . The reaction is:



If the volume of the 3.00-liter flask at 1000°C when a concentration of H_2 is present at equilibrium:

- What is the concentration of H_2 at equilibrium?
- What is the concentration of S_2 at equilibrium?
- What is the concentration of H_2S at equilibrium?
- What is the value of K_c at equilibrium?
- What other equilibria have chemical formulas involving H_2 and S_2 at equilibrium above?
- Calculate:

	H_2	S_2	H_2S
Initial (moles)			
Change (moles)			
Final (moles)	0.020	0.010	

7. (20) Equilibrium constants:

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Table of Contents Answers For Sch4u Unit 4

1. Understanding the eBook Answers For Sch4u Unit 4
 - The Rise of Digital Reading Answers For Sch4u Unit 4
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Sch4u Unit 4
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Sch4u Unit 4
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Sch4u Unit 4
 - Personalized Recommendations

- Answers For Sch4u Unit 4 User Reviews and Ratings
- Answers For Sch4u Unit 4 and Bestseller Lists
- 5. Accessing Answers For Sch4u Unit 4 Free and Paid eBooks
 - Answers For Sch4u Unit 4 Public Domain eBooks
 - Answers For Sch4u Unit 4 eBook Subscription Services
 - Answers For Sch4u Unit 4 Budget-Friendly Options
- 6. Navigating Answers For Sch4u Unit 4 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Sch4u Unit 4 Compatibility with Devices
 - Answers For Sch4u Unit 4 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Sch4u Unit 4
 - Highlighting and Note-Taking Answers For Sch4u Unit 4
 - Interactive Elements Answers For Sch4u Unit 4
- 8. Staying Engaged with Answers For Sch4u Unit 4
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Sch4u Unit 4
- 9. Balancing eBooks and Physical Books Answers For Sch4u Unit 4
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Sch4u Unit 4
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For Sch4u Unit 4
 - Setting Reading Goals Answers For Sch4u Unit 4
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers For Sch4u Unit 4
 - Fact-Checking eBook Content of Answers For Sch4u Unit 4

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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