

**Review for the Full Year General Chemistry
American Chemical Society Exam**

(a) Thermodynamics, (b) Transition Metals, (c) Redox and Electrochemistry, (d) Nuclear and (e) Organic

THERMODYNAMICS

84. A 10.0 g sample of silver is heated to 100.0 °C and then added to 20.0 g of water at 23.0 °C in an insulated calorimeter. At thermal equilibrium the temperature of the system was measured as 25.0 °C. What is the specific heat of silver? [specific heat = $4.2 \text{ J g}^{-1} \text{ }^{\circ}\text{C}^{-1}$]

- (A) 0.11 J/g K (C) 17 J/g K
(B) 0.22 J/g K (D) 34 J/g K

85. What is ΔH in kJ for the reaction:



- (A) -1830 kJ (B) -2284 kJ
(C) -2510 kJ (D) -1605 kJ

86. Given these equations:



calculate the heat of formation of $\text{SO}_3(\text{g})$.

- (A) -500 kJ mol⁻¹ (C) +100 kJ mol⁻¹
(B) -400 kJ mol⁻¹ (D) +200 kJ mol⁻¹

87. Given:



How much energy is produced in the combustion of 28 g of CO? [Atomic Masses: C 12.0 g mol⁻¹; O 16.0 g mol⁻¹]

- (A) 88 kJ (B) 109 kJ
(C) 175 kJ (D) 284 kJ

88. Calculate the value of ΔH (in kJ mol⁻¹) for the reaction:



Bond Energies (kJ mol ⁻¹)	
H-H	435
N=N	945 (in N ₂)
N-H	389

- (A) 2340 kJ of heat absorbed
(B) 213 kJ of heat absorbed
(C) 2340 kJ of heat evolved
(D) 83 kJ of heat evolved

89. Which change is likely to be accompanied by the greatest increase in entropy?



90. For which process is the entropy change per mole the largest at constant temperature?



91. In which process is entropy decreased?

- (A) dissolving sugar in water
(B) expanding a gas
(C) evaporating a liquid
(D) freezing water

92. When $\text{Al}_2\text{O}_3(\text{s})$ is formed from the elements at standard conditions, the values of ΔH° and ΔG° at 298 K are -1676 kJ mol⁻¹ and -1577 kJ mol⁻¹, respectively. The standard entropy of formation per mole, in joules per degree, will be

- (A) -332 (B) -157 (C) -93.3
(D) -0.0933 (E) +15.7

93. Vaporization of a liquid is an example of a process for which

- (A) ΔH , ΔS , and ΔG are positive at all temperatures.
(B) ΔH and ΔS are positive.
(C) ΔG is negative at low temperatures, positive at high temperatures.
(D) $\Delta H = \Delta S$

94. A particular chemical reaction has a negative ΔH and negative ΔS . Which statement is correct?

- (A) The reaction is spontaneous at all temperatures.
(B) The reaction is nonspontaneous at all temperatures.
(C) The reaction becomes spontaneous as temperature increases.
(D) The reaction becomes spontaneous as temperature decreases.

American Chemical Society Exam

I. Dwaine Eubanks



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The Fractured Marketplace for Standardized Testing Walter M. Haney, George F. Madaus, Robert Lyons, 2012-12-06 Standardized testing in the United States has been increasing at a rapid pace in the last twenty five years The market for tests has not only been expanding rapidly but has also been changing sharply in structure into a fractured marketplace Indeed one of the main features of this book is that the market for standardized testing is highly fractured with

segments of the market facing monopoly conditions others facing oligopoly conditions and still others where near free market conditions exist One of the main premises of the book is that the structures of markets have strong implications for how those markets perform While this notion is widely accepted among economists it is not widely appreciated in educational research A second motivation for the book is that very little scholarly attention has been focused on the standardized testing industry This topic the structure of the testing industry and implications for the quality of tests and test use affects how we evaluate the learning of students the effectiveness of teaching the quality of schools and the educational health of the nation Of particular concern to the authors is one vital aspect of test quality test validity This book is the most current and authoritative review and analysis of the market for standardized testing *Learning and Understanding* National Research Council, Division of Behavioral and Social Sciences and Education, Center for Education, Committee on Programs for Advanced Study of Mathematics and Science in American High Schools, 2002-08-06 This book takes a fresh look at programs for advanced studies for high school students in the United States with a particular focus on the Advanced Placement and the International Baccalaureate programs and asks how advanced studies can be significantly improved in general It also examines two of the core issues surrounding these programs they can have a profound impact on other components of the education system and participation in the programs has become key to admission at selective institutions of higher education By looking at what could enhance the quality of high school advanced study programs as well as what precedes and comes after these programs this report provides teachers parents curriculum developers administrators college science and mathematics faculty and the educational research community with a detailed assessment that can be used to guide change within advanced study programs

Confrontations Don Williamson, 2020-05-15 Book Delisted

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