

Name: _____

Date: _____

Civil War Study Guide Answer Key

1. Life was hard in the South on the home front because _____. Prices were high and there was not enough food.
2. Sherman's strategy of total war included _____. Attacking and destroying anything the enemy could use to continue fighting.
3. After the Civil War, President Lincoln wanted _____. Defeated Confederate states to set up new governments and rejoin the Union quickly.
4. At the end of Reconstruction:
True or False: African Americans lost political power.
True or False: The president told federal troops to leave the South.
True or False: Segregation was enforced.
5. Why did Jim Crow laws harm African Americans? _____. They enforced the segregation of African Americans from other Americans.
6. What are some advantages the North had during the Civil War? _____. The North had more factories for making weapons and supplies.
7. What took the place of slavery? _____. Sharecropping.
8. The Union was also known as the _____. The North and The Yankees.
9. Why did Southern states leave the Union _____. To protect their own right to own slaves.
10. Who won the battle at Gettysburg? _____. the Union.
11. Before the Civil War, why did tension increase between the northern and southern states? _____. They disagreed about whether or not slavery should continue.
12. Who was the president of the Union during the Civil War? _____. Abraham Lincoln.
13. Who led the Confederate troops at the Battle of Bull Run? _____. Thomas "Stonewall" Jackson.
14. What is the name of the famous speech Lincoln gave during the Civil War? _____. Gettysburg Address.
15. What was the Underground Railroad? _____. a system of escape trails that led slaves to freedom.
16. The battles of the Civil War were mostly fought _____. in the South.
17. What Constitutional Amendment ended slavery in the United States? _____. 13th.

Answer Key To The Civil War Begins Reteach

J Ma



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McGraw-Hill Reading, 200? World History: Connections to Today Elisabeth Gaynor Ellis, Pearson/Prentice Hall, 2003-12 **Houghton Mifflin English** Houghton Mifflin, Houghton Mifflin Reading David J. Cooper, 2001 Defamiliarizing Japan's Asia-Pacific War W. Puck Brecher, Michael W. Myers, 2019-10-31 This wide ranging collection seeks to reassess conventional understanding of Japan's Asia Pacific War by defamiliarizing and expanding the rhetorical narrative Its nine chapters diverse in theme and method are united in their goal to recover a measured historicity about the conflict by either introducing new areas of knowledge or reinterpreting existing ones Collectively they cast doubt on the war as familiar and recognizable compelling readers to view it with fresh eyes Following an introduction that problematizes timeworn narratives about a unified Japan and its illegal war or race war early chapters on the destruction of Japan's diplomatic records and government interest in an egalitarian health care policy before during and after the war oblige us to question selective histories and moral judgments about wartime Japan The discussion then turns to artistic cultural production and self determination specifically to Osaka rakugo performers who used comedy to contend with state oppression and to the role of women in creating care packages for soldiers abroad Other chapters cast doubt on well trod stereotypes Japan's lack of pragmatism in its diplomatic relations with neutral nations and its irrational and fatalistic military leadership and examine resistance to the war by a prominent Japanese Christian intellectual The volume concludes with two nuanced responses to race in wartime Japan one maintaining the importance of racial categories while recognizing the performance of Japaneseness the other observing that communities often reflected official government policies through nationality rather than race Contrasting findings like these underscore

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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should look to the appropriate text section to see how they are treated pure acid 0 mL of base is added section 17.3.2.1

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web solution a titrant volume 0 mL the solution pH is due to the acid ionization of HCl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the pH of the solution is then $pH = -\log 0.100 = 1.000$ pH $\log 0.100 = 1.000$ b titrant volume 12.50 mL

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web at the equivalence point in an acid base titration moles of base = moles of acid and the solution only contains salt and water acid base titrations are monitored by the change of pH as titration progresses

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web expert answer 100% 2 ratings the example is based upon Henderson-Hasselbalch equation the equation gives relation between pH and pK_a of weak acid the equation is $pH = pK_a + \log \frac{[salt]}{[acid]}$ at equivalence point all a view the full answer transcribed image text

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web ab acid base titration curves purpose the purpose of this lab is to determine the shape of acid base titration curves theoretically and experimentally and to use an acid base titration to measure the effectiveness of antacid tablets summary a titration curve is the plot of the pH value of a solution vs the amount of titrant added

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web pH titration graph 1 lab data provide columns a and b 2 a graph of pH column b plotted as ordinate vs volume column a as abscissa provides a titration curve as illustrated by graph 1 or 2 below 3 estimate of the midpoint of the steeply rising section of the titration curve provides an evaluation of the equivalence point

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web acid base titration lab report jessica chung chem 267 lab date completed 25 september 2019 date turned in 9 october 2019 figure 1 acid base titration curve from titration with indicator for part 1 volume of NaOH added mL pH of the solution

Volume of NaOH added (mL)	pH of the solution
4	6.4
6	4.6
4	6.5
5	9.5
5	9.5
9	4.6

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web background plot of the pH of a solution against the volume of titrant added is called a titration curve the pH can be measured directly with a pH meter while titrant is added from a buret for the acids used in this experiment the titrant will always be a 0.1 M solution of the strong base NaOH

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web oct 30 2022 there are two basic types of acid base titrations indicator and potentiometric in an indicator based titration you add another chemical that changes color at the ph equal to the equivalence point when the acid and base are in
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web rp 1 make up a volumetric solution and carry out a simple acid base titration physical chemistry acids and bases ph curves titrations and indicators titrations of acids with bases students should be able to perform calculations for these titrations based on experimental results inorganic chemistry transition metals variable

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web strong acids and strong bases completely ionize in solution resulting in water and a salt for a weak acid there s only partial ionization the titration curve of a strong base strong acid showed a stretched out curve as it started with a slow gradual change in ph as it reached the equivalence point