

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive the structural formula of a given chemical compound without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from weighing information in the lab.
3. To determine the molar mass of a solid.
4. To determine the stoichiometric amount of a solid compound that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction Information:

Reaction of Sodium Chloride and Silver Nitrate: $2\text{AgNO}_3 + 2\text{NaCl} \rightarrow 2\text{AgCl} + 2\text{NaNO}_3$

Reaction of Silver Nitrate and Sodium Chloride: $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$

Reaction of Silver Nitrate and Sodium Chloride: $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$

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Procedure:

Reaction and Calculations:

1. Reaction:

$$\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$$

$$\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$$

$$\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$$

$$\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$$

2. Reaction:

Reaction:

$$\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$$

$$\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$$

$$\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$$

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