

The Advanced Placement Examination in Chemistry

Part II - Free Response Questions & Answers 1970 to 2008 Thermodynamics

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Questions With ANSWERS

1979 B

Compound	ΔH_f° (kilocalories/mole)	S° (calories/mole K)
$H_2O(l)$	-68.3	16.7
$CO_2(g)$	-94.1	51.1
$O_2(g)$	0.0	49.0
C_3H_8	?	64.5

When 1.000 gram of propane gas, C_3H_8 , is burned at 25°C and 1.00 atmosphere, $H_2O(l)$ and $CO_2(g)$ are formed with the evolution of 12.03 kilocalories.

- Write a balanced equation for the combustion reaction.
- Calculate the molar enthalpy of combustion, ΔH_{comb}° , of propane.
- Calculate the standard molar enthalpy of formation, ΔH_f° , of propane gas.
- Calculate the entropy change, ΔS_{comb}° , for the reaction and account for the sign ΔS_{comb}° .

Answer:



(b) $\Delta H_{comb}^\circ = \frac{-12.03 \text{ kcal}}{1.000 \text{ g}} \times \frac{44.10 \text{ g}}{1 \text{ mol}} = -530.5 \text{ kcal/mol}$

This answer should be **-529.3 kcal/mol**

(c) $\Delta H_{comb}^\circ = 3\Delta H_f^\circ CO_2 + 4\Delta H_f^\circ H_2O - \Delta H_f^\circ C_3H_8$
 $-529.3 \text{ kcal} = [3(-94.1) + 4(-68.3) - X] \text{ kcal}$
 $\Delta H_f^\circ = -26.2 \text{ kcal/mol}$

(d) $\Delta S_{comb}^\circ = [3S_{CO_2} + 4S_{H_2O}] - [S_{C_3H_8} + 5S_{O_2}] = [3(51.1) + 4(16.7)] - [64.5 + 5(49.0)]$
 $= -89.4 \text{ cal/mol}\cdot\text{K}$

Entropy decreases due to loss of highly disordered gaseous species upon combustion.

2008 Ap Chemistry Answers

TD Snyder



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