

8. Sulfur, having atomic no. 16, is found within the Earth's crust. Sulfur is composed out of isotopes of almost all existing isotopes.

18 isotopes of sulfur from 18 isotopes were employed to give the following composition of isotopes.

isotope	abundance (%)
^{32}S	95.02
^{34}S	4.29
^{36}S	0.69

- (a) Define the term relative atomic mass.

The relative atomic mass of an element is the average mass of its atoms, measured on the scale of ^{12}C , which is taken as 12.

- (b) Calculate the relative atomic mass of the isotopes of sulfur.

They give values in three decimal places.

$$\left(\frac{32 \times 95.02}{100} \right) + \left(\frac{34 \times 4.29}{100} \right) + \left(\frac{36 \times 0.69}{100} \right)$$

$$= 32.07$$

$$\text{Answer} = 32.07 \text{ amu}$$

- (c) Some isotopes are easily found, naturally occurring, isotopes that artificially were made up of any particles which could be used for the purpose, isotopes known as the products of subatomic particles in strong and in ions.

Complete the table to show the number of subatomic particles in the ^{32}S atom and $^{32}\text{S}^{2-}$ ion.

	protons	neutrons	electrons
^{32}S	16	16	16
$^{32}\text{S}^{2-}$	16	16	18



Answer To F321 Chemistry May Series

Douglas Buchanan, James Melrose



Answer To F321 Chemistry May Series:

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