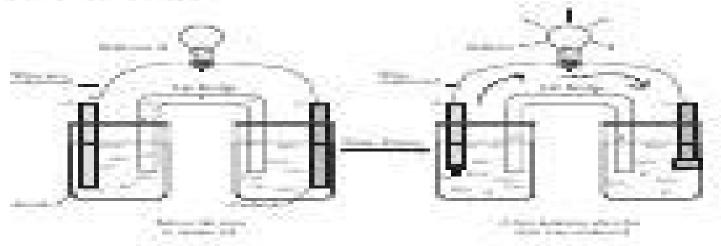
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applies only to electrolysis passing electricity through a chemical to split it up Batteries are like electrolysis going backwards they split up chemicals to make electricity so the terms anode and cathode are switched around Okay To avoid confusion I suggest it s best not to use the terms anode and cathode at all It s better to say positive terminal and negative terminal and then it s always clear what you mean whether you re talking about batteries or electrolysis Chemical reactions Now back to our battery The positive and negative electrodes are separated by the chemical electrolyte It can be a liquid but in an ordinary battery it is more likely to be a dry powder When you connect the battery to a lamp and switch on chemical reactions start happening One of the reactions generates positive ions shown here as big yellow blobs and electrons smaller brown blobs at the negative electrode The positive ions flow through the electrolyte to the positive electrode from the green line to the red one Meanwhile the electrons smaller brown blobs flow around the outside circuit blue line to the positive electrode and make the lamp light up on the way The electrons and ions flow because of the chemical reactions happening inside the battery usually two or three of them going on simultaneously The exact reactions depend on the materials from which the electrodes and electrolyte are made and we won t go into them here If you want to know what they are Google the type of the battery you re interested in followed by the words anode cathode reactions For example Google zinc carbon battery anode cathode reactions Whatever chemical reactions take place the general principle of electrons going around the outer circuit and ions flowing in the opposite direction through the electrolyte happens in all batteries As the battery generates power the chemicals inside it are gradually converted into different chemicals

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