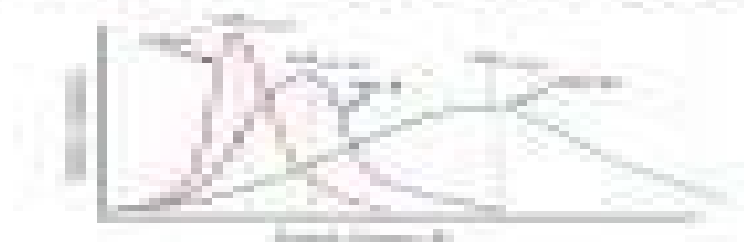


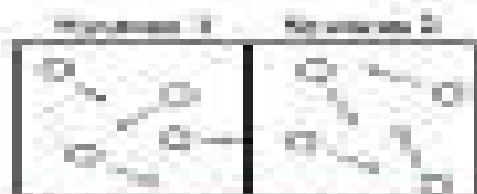
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with m in cm^2/s^2 is the mass of particles and molecules mass and change, directly connected as all of temperature increase. The Boltzmann distribution curve shows that the average velocity of all of the distribution's molecules (area under the curve) is proportional to temperature.

3. Sketch a curve to represent the relative distribution of the particles in the following two gas samples at the same gas:



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.

System 1 is at a higher temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the gas particles are at a higher KE, and all the individual particles have the same KE. According to the equation $KE = \frac{1}{2}mv^2$, all of the particles have the same mass, so the only way to get a higher KE is to have a higher velocity. In System 2, the particles are at a lower KE, and all the individual particles have the same KE. According to the equation $KE = \frac{1}{2}mv^2$, all of the particles have the same mass, so the only way to get a lower KE is to have a lower velocity.

- b. Are there different total kinetic energy for System 1? Justify your answer in terms of molecular collisions.

Yes, there is a difference in total kinetic energy. In System 1, the gas particles are at a higher KE, and all the individual particles have the same KE. According to the equation $KE = \frac{1}{2}mv^2$, all of the particles have the same mass, so the only way to get a higher KE is to have a higher velocity. In System 2, the gas particles are at a lower KE, and all the individual particles have the same KE. According to the equation $KE = \frac{1}{2}mv^2$, all of the particles have the same mass, so the only way to get a lower KE is to have a lower velocity.

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