

6. Given the following standard reduction of a galvanic cell:



- a. Identify the anode and cathode half-reactions:



Note: This problem is an *oxidation-reduction* problem, not an *acid-base* problem.

- b. Write the overall balanced cell reaction:



7. Write the cell notation for a galvanic cell that uses the reactions below:



8. A galvanic cell is constructed based on the following reaction:



- a. Calculate the standard free energy change (ΔG°) for this reaction at 298 K.



- b. Is this reaction spontaneous?

Yes, this reaction is spontaneous at standard state conditions, as indicated by the negative sign of ΔG° .

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