

**Unit four: Principles of supporting change in a business environment****Assessment**

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**Please note that this Assessment document has 3 pages and is made up of 3 Sections.**

**Name: D S**

**Section 1: Why change happens**

This section will help you to evidence Learning Outcome 1: **Understand why change happens in a business environment.**

| Learning objective                                       | Place in Assessment      |
|----------------------------------------------------------|--------------------------|
| 1.1 Explain reasons for change in a business environment | <b>Question 1 Page 1</b> |

**1. Explain why change happens in a business environment. You should include at least three reasons in your answer. [1.1]**

In today's competitive world, change is inevitable, essential and natural, whether it is for people or businesses. Without it, no business can flourish or retain its position for a long time. Changes can be driven externally or internally.

External changes happen outside a business environment. They are changes in government, needs of the consumers, competition to a business and introduction of new technological trends. Analysis of the above help a business prepare and respond to these waves of modifications. This is called a PEST analysis that covers all the aspects mentioned above, political, economic, social or technological.

Customers or consumers is a social force of any business. They get affected by any impact on the economy, variations in their own lifestyles or what is happening around the world, therefore, it becomes essential for any business to offer deals which not only helps to retain existing customers but also make new ones.

Similarly, changes in the government policies, laws, interest rates can directly affect a company's bottom line, especially if they operate in other countries, as they have to then adjust according to fluctuations in that economy.

Technology advancement is a continual process and every business has to be nimble to keep up with the changing technological trends. The most common example of the same is online

# Answers For Unit 107 Business Admin

**Jeffrey H. Dyer, Paul Godfrey, Robert  
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web jan 29 2020 solution number of moles of solute ethyl alcohol  $n_b$  23 g 46 g mol 1 0 5 mol number of moles of solvent water  $n_a$  54 g 18 g mol 1 3 mol total number of moles  $n_a + n_b$  0 5 3 3 5 mol mole fraction of solute ethyl alcohol  $x_b = \frac{n_b}{n_a + n_b}$

b 0 5 3 5 0 1429

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web question 1 find the mole fraction of methanol  $\text{CH}_3\text{OH}$  and water in a solution prepared by dissolving 4.5 g of alcohol in 40 g of  $\text{H}_2\text{O}$  molar mass of  $\text{H}_2\text{O}$  is 18 gm/mole and molar mass of  $\text{CH}_3\text{OH}$  is 32 gm/mole solution moles of  $\text{CH}_3\text{OH}$  4.5/32 = 0.14 mole moles of  $\text{H}_2\text{O}$  40/18 = 2.2 moles therefore according to the equation

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web we used mole fractions in gen chem 1 to describe gas mixtures where the partial pressure of a gas was its mole fraction time the total pressure  $p_a = x_a \cdot p_{\text{total}}$  mole fraction  $x_a = \frac{\text{moles of substance a}}{\text{total moles}}$  mole fraction of substance a is often denoted by  $x_a$  and

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web 1 in the given mixture the sum of all the mole fractions is equal to one  $\sum x_i = 1$  the multiplication of the mole fraction by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction mole fraction is not dependent on the temperature

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web nov 6 2023  $x_a \cdot \text{mol a} + x_b \cdot \text{mol b} = 1$  note in the given mixture the sum of all the mole fractions present is equal to one  $\sum x_i = 1$  when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is a unitless and dimensionless expression image will be uploaded soon

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web solution total mass of the answer 100 g mass of benzene 30 g therefore mass of carbon tet 100 - 30 = 70 g benzene s molar mass  $\text{C}_6\text{H}_6$  12.011  $\times$  6 + 1.008  $\times$  6 = 78.11 g/mol hence  $\text{C}_6\text{H}_6$  s number of moles 30/78.11 = 0.3846 mol molar mass of carbon tet  $\text{CCl}_4$  12.011  $\times$  1 + 35.453  $\times$  4 = 153.81 g/mol

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web jan 30 2020 ans the molality of solution is 1 mol/kg or 1 m example 0.211 g of urea  $\text{NH}_2\text{CONH}_2$  was dissolved in 100 g of water calculate the molarity and molality of the solution given n 14 h 1 c 12 o 16 given mass of solute urea 11.11 g mass

of solvent water 100 g 0 1 kg to find molarity of solution solution

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web solution moles of  $\text{CH}_3\text{OH}$  5 5 32 0 17 mole moles of  $\text{H}_2\text{O}$  40 18 2 2 moles therefore according to the equation mole fraction of  $\text{CH}_3\text{OH}$  0 17 2 2 0 17 mole fraction of  $\text{CH}_3\text{OH}$  0 073 to solve more examples on mole fraction formula and practice more question please visit byju s com

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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0 449 molal  $\text{KBr}$  that contains 2 92 kg of water 1 molality moles solute divided by kilograms solute 0 449 mol kg x 2 92 kg x 1 31108 mol of  $\text{KBr}$  2 moles times molar mass equals grams 1 31108 mol 119 0023 g mol 156 g  $\text{KBr}$

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web calculate the mass percent mole fraction molality and molarity of hexane  $\text{C}_6\text{H}_{14}$  d 0 660 g  $\text{cm}^3$  when it is mixed with heptane  $\text{C}_7\text{H}_{16}$  d 0 684 g  $\text{cm}^3$  in a 35 0 55 0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web since the mole fraction of the o 2 is 0 050 it contributes 0 050 or if you will 5 0 of the total pressure thus total pressure is 0 21 0 050 4 2 atm problem 9 a sample of 1 43 g of helium and an unweighed quantity of o 2 are mixed in a flask at room temperature