

9) A 250mL solution of 0.5M NaF has 50mL of a 1M BaF₂ solution added to it. What is the concentration of [F⁻] in the new solution? Assume that volume is additive.

From (8)

$$[F^-] = 0.125 \text{ mol} \left| \begin{array}{l} 0.05 \text{ L} \times 1 \text{ M BaF}_2 \\ = 0.05 \text{ mol BaF}_2 \\ = 0.1 \text{ mol F}^- \end{array} \right| \begin{array}{l} \text{Final sol'n} \\ \text{mol F}^- = 0.125 \text{ mol} + 0.1 \text{ mol} = 0.225 \text{ mol F}^- \\ \text{volume} = 0.25 \text{ L} + 0.05 \text{ L} \\ = 0.3 \text{ L} \end{array}$$

$$M = \frac{0.225 \text{ mol}}{0.3 \text{ L}} = \boxed{0.75 \text{ M}}$$

10) How many grams of NaHCO₃ (molar mass = 84.01 g/mol) are needed to make a 350mL solution of 0.75M NaHCO₃?

$$n = V \times M \\ = 0.35 \text{ L} \times 0.75 \text{ M NaHCO}_3 \\ = 0.2625 \text{ mol NaHCO}_3$$

$$\text{mass} = 0.2625 \text{ mol NaHCO}_3 \times \frac{84.01 \text{ g NaHCO}_3}{1 \text{ mol NaHCO}_3} \\ = \boxed{22.05 \text{ g NaHCO}_3}$$

11) Calculate the amount of distilled water that is needed to dilute a 3L solution of 5M HBr to a concentration of 1.5M HBr?

$$\text{mol} = 3 \text{ L} \times 5 \text{ M} \\ = 15 \text{ mol HBr}$$

$$\text{volume needed: } V = \frac{15 \text{ mol}}{1.5 \text{ M HBr}} \\ = 10 \text{ L}$$

$$10 \text{ L} - 3 \text{ L} = \boxed{7 \text{ L needed to add}}$$

12) What volume of 8M H₂SO₄ is needed to prepare a 100mL solution of 0.5M H₂SO₄?

$$0.1 \text{ L} \times 0.5 \text{ M H}_2\text{SO}_4 \\ = 0.005 \text{ mol H}_2\text{SO}_4 \text{ needed}$$

$$V_{\text{ol}} = \frac{0.005 \text{ mol H}_2\text{SO}_4}{8 \text{ M H}_2\text{SO}_4}$$

$$= 6.25 \times 10^{-4} \text{ L}$$

$$= \boxed{0.625 \text{ mL of } 8 \text{ M H}_2\text{SO}_4, \text{ the rest is water}}$$

13) The maximum amount of salt that can be contained in water is approximately 26.3% by weight. Convert this into concentration with units of molarity. Info: 1g of water = 1 mL of water.

$$\text{assume } 100 \text{ g} \\ 26.3 \text{ g NaCl}$$

$$26.3 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.44 \text{ g NaCl}} = 0.45 \text{ mol NaCl}$$

$$73.7 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18 \text{ g H}_2\text{O}} = 4.1 \text{ mol H}_2\text{O} / \text{sol'n}$$

$$M = \frac{0.45 \text{ NaCl}}{0.0737 \text{ L sol'n}} = \boxed{6.1 \text{ M}}$$

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Table of Contents All Pogil Chemistry Answers Molarity

1. Understanding the eBook All Pogil Chemistry Answers Molarity
 - The Rise of Digital Reading All Pogil Chemistry Answers Molarity
 - Advantages of eBooks Over Traditional Books
2. Identifying All Pogil Chemistry Answers Molarity
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an All Pogil Chemistry Answers Molarity
 - User-Friendly Interface
4. Exploring eBook Recommendations from All Pogil Chemistry Answers Molarity
 - Personalized Recommendations

- All Pogil Chemistry Answers Molarity User Reviews and Ratings
- All Pogil Chemistry Answers Molarity and Bestseller Lists
- 5. Accessing All Pogil Chemistry Answers Molarity Free and Paid eBooks
 - All Pogil Chemistry Answers Molarity Public Domain eBooks
 - All Pogil Chemistry Answers Molarity eBook Subscription Services
 - All Pogil Chemistry Answers Molarity Budget-Friendly Options
- 6. Navigating All Pogil Chemistry Answers Molarity eBook Formats
 - ePub, PDF, MOBI, and More
 - All Pogil Chemistry Answers Molarity Compatibility with Devices
 - All Pogil Chemistry Answers Molarity Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of All Pogil Chemistry Answers Molarity
 - Highlighting and Note-Taking All Pogil Chemistry Answers Molarity
 - Interactive Elements All Pogil Chemistry Answers Molarity
- 8. Staying Engaged with All Pogil Chemistry Answers Molarity
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers All Pogil Chemistry Answers Molarity
- 9. Balancing eBooks and Physical Books All Pogil Chemistry Answers Molarity
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection All Pogil Chemistry Answers Molarity
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine All Pogil Chemistry Answers Molarity
 - Setting Reading Goals All Pogil Chemistry Answers Molarity
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of All Pogil Chemistry Answers Molarity
 - Fact-Checking eBook Content of All Pogil Chemistry Answers Molarity

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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