

## CHAPTER 7

## STUDY GUIDE

## Ionic Compounds and Metals

### Section 7.1 Ion Formation

*In your textbook, read about chemical bonds and formation of ions.*

Use each of the terms below just once to complete the passage.

|               |           |                             |         |             |
|---------------|-----------|-----------------------------|---------|-------------|
| chemical bond | electrons | energy level                | ions    | noble gases |
| nucleus       | octet     | pseudo-noble gas formations | valence |             |

The force that holds two atoms together is called a(n) (1) \_\_\_\_\_.

Such an attachment may form by the attraction of the positively charged

(2) \_\_\_\_\_ of one atom for the negatively charged

(3) \_\_\_\_\_ of another atom, or by the attraction of charged atoms, which are called

(4) \_\_\_\_\_. The attractions may also involve

(5) \_\_\_\_\_ electrons, which are the electrons in the outermost

(6) \_\_\_\_\_. The (7) \_\_\_\_\_ are a family of elements that have very

little tendency to react. Most of these elements have a set of eight outermost electrons, which is called a

stable (8) \_\_\_\_\_. The relatively stable electron structures developed by loss of

electrons in certain elements of groups 3, 4, 13, and 14 are called (9) \_\_\_\_\_.

For each statement below, write *true* or *false*.

- \_\_\_\_\_ 10. A positively charged ion is called an anion.
- \_\_\_\_\_ 11. Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge.
- \_\_\_\_\_ 12. Elements tend to react so that they acquire the electron structure of a halogen.
- \_\_\_\_\_ 13. A sodium atom tends to lose one electron when it reacts.
- \_\_\_\_\_ 14. The electron structure of a zinc ion ( $\text{Zn}^{2+}$ ) is an example of a pseudo-noble gas formation.
- \_\_\_\_\_ 15. A  $\text{Cl}^-$  ion is an example of a cation.
- \_\_\_\_\_ 16. The ending *-ide* is used to designate an anion.
- \_\_\_\_\_ 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions.

# Answer Study Guide Ionic Compounds And Metals

**David Baud**



## **Answer Study Guide Ionic Compounds And Metals:**

**AP Biology Study Guide** Sundar Nathan, 2009-11 Sundar Nathan received a Bachelor's degree in Electrical Engineering from Anna University Chennai India and a Masters degree in Biomedical Engineering from the University of Texas at Austin Working for over a year with a team of talented Phds MPhils and MScs from all over the world Sundar compiled this comprehensive study guide to help students prepare diligently understand the concepts and Crush the AP Bio Test **Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter **"O" Level Study Guide - Chemistry Quite Easily Done ,**

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takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

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**Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today** Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text

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