

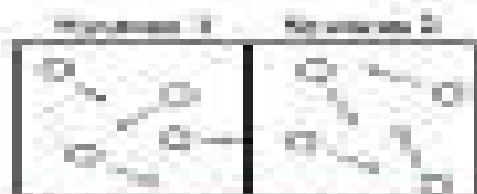
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The faster molecules are moving, meaning that in terms of the rate, more of them are moving faster and changing, resulting in a higher rate of temperature increase. The Boltzmann distribution curve shows that the average velocity of each of the distributions is increasing from A to C, and the average temperature is increasing.

3. Sketch a curve used to represent the relative distribution of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are more spread out, and the bonds between them are weaker, indicating a lower temperature. In System 2, the particles are more spread out, and the bonds between them are stronger, indicating a higher temperature.
- b. An initial amount of System 1 energy for System 2. Justify your answer in terms of molecular collisions.
The initial energy of System 1 is higher than the initial energy of System 2. The particles in System 1 are more spread out, and the bonds between them are weaker, indicating a lower temperature. The particles in System 2 are more spread out, and the bonds between them are stronger, indicating a higher temperature.

Answer Key Unit 6 Thermochemistry

J Ma



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