

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product produced.
5. To compare the theoretical amount to the actual amount of product formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Procedure:

Calculations and Results:

1. Calculations:

$$\text{moles of sodium chloride} = \frac{\text{mass of NaCl}}{\text{molar mass of NaCl}} = \frac{5.00 \text{ g}}{58.44 \text{ g/mol}} = 0.0855 \text{ mol}$$

$$\text{moles of calcium chloride} = \frac{\text{mass of CaCl}_2}{\text{molar mass of CaCl}_2} = \frac{5.00 \text{ g}}{110.98 \text{ g/mol}} = 0.0450 \text{ mol}$$

$$\text{moles of sodium hydroxide produced} = \frac{\text{mass of NaOH}}{\text{molar mass of NaOH}} = \frac{5.00 \text{ g}}{40.00 \text{ g/mol}} = 0.125 \text{ mol}$$

$$\text{moles of product sodium chloride} = \frac{\text{mass of NaCl}}{\text{molar mass of NaCl}} = \frac{5.00 \text{ g}}{58.44 \text{ g/mol}} = 0.0855 \text{ mol}$$

2. Chemical Equations: $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$

3. Reaction:

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

$$\text{moles of NaCl} = \frac{\text{mass of NaCl}}{\text{molar mass of NaCl}} = \frac{5.00 \text{ g}}{58.44 \text{ g/mol}} = 0.0855 \text{ mol}$$

Answers For Lab 6b Heath Chemistry

Rachel S Tattersall



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1. Understanding the eBook Answers For Lab 6b Heath Chemistry
 - The Rise of Digital Reading Answers For Lab 6b Heath Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Lab 6b Heath Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Lab 6b Heath Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Lab 6b Heath Chemistry
 - Personalized Recommendations
 - Answers For Lab 6b Heath Chemistry User Reviews and Ratings
 - Answers For Lab 6b Heath Chemistry and Bestseller Lists
5. Accessing Answers For Lab 6b Heath Chemistry Free and Paid eBooks

- Answers For Lab 6b Heath Chemistry Public Domain eBooks
- Answers For Lab 6b Heath Chemistry eBook Subscription Services
- Answers For Lab 6b Heath Chemistry Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Answers For Lab 6b Heath Chemistry Compatibility with Devices
 - Answers For Lab 6b Heath Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Lab 6b Heath Chemistry
 - Highlighting and Note-Taking Answers For Lab 6b Heath Chemistry
 - Interactive Elements Answers For Lab 6b Heath Chemistry
- 8. Staying Engaged with Answers For Lab 6b Heath Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Lab 6b Heath Chemistry
- 9. Balancing eBooks and Physical Books Answers For Lab 6b Heath Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Lab 6b Heath Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For Lab 6b Heath Chemistry
 - Setting Reading Goals Answers For Lab 6b Heath Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers For Lab 6b Heath Chemistry
 - Fact-Checking eBook Content of Answers For Lab 6b Heath Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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