

# অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পটাসিয়াম কপ্টার, এখানে মোট পটাসিয়াম, আরও কপ্টারের শীর্ষস্থানীয় কপ্টারের পটাসিয়াম এবং মনোমের টেক্টর পটাসিয়াম  
ন্যূনতম উত্তর অধ্যায়ভিত্তিক মেট্রিক্স রয়েছে। এগুলোর অনুশীলন করলে তুমি এ অধ্যায়টি থেকে সেরাভাবে বুঝতে পারবে।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০।এস. ১।

- মোলের প্রকণ কী?
- "কেলস পানিদ্রুত লবণের প্রাকৃত্য, 'অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়'— ব্যাখ্যা করো।
- A-পাত্রে বিন্যাসন এলিটের ভর কত নির্ণয় করো।
- প্রকণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা পরিণত হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL প্রকণে কোন প্রকণের 1 mol প্রকণিত থাকলে উক্ত প্রকণকে মোলার প্রকণ বলে।

কেলস পানিদ্রুত লবণের প্রাকৃত্য অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, আপমাত্রা বৃদ্ধির সাথে সাথে প্রাকৃত্য বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট আপমাত্রার কেলস পানিদ্রুত লবণ নিষ্কৃতি হয়। তখন প্রাকৃত্য হ্রাস পেতে থাকে। উদাহরণস্বরূপ  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর কথা বলা যায়।  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর ক্ষেত্রে পানিতে প্রকণিত করার সময় বিয়োজন প্রক্রিয়াটি অপেক্ষারী প্রক্রিয়া কিন্তু  $32.4^\circ\text{C}$  আপমাত্রার উপরে  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  যখন  $\text{Na}_2\text{SO}_4$  এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি অপেক্ষারী হয়। এ কারণে  $32.4^\circ\text{C}$  আপমাত্রার উপরে আপমাত্রা বৃদ্ধিজনিত কারণে প্রাকৃত্য হ্রাস ঘটে।

২

আমরা জানি,  
 $w = MV$   
 $= (0.5 \times 98 \times 0.25) \text{ g}$   
 $= 12.25 \text{ g}$

অতএব, প্রকণিত  $\text{H}_2\text{SO}_4$  এর ভর  
 $= 12.25 \text{ g}$

উদাহরণের A-পাত্রে অবস্থিত,  
 $\text{H}_2\text{SO}_4$  এর আণবিক ভর,  
 $M = 1 \times 2 + 32 + 16 \times 4$   
 $= 2 + 32 + 64$   
 $= 98 \text{ g/mol}$

প্রকণের ঘনমাত্রা,  
 $S = 0.5 \text{ M}$   
প্রকণের আয়তন,  
 $V = 250 \text{ mL}$   
 $= \frac{250}{1000} = 0.25 \text{ L}$

প্রকণিত  $\text{H}_2\text{SO}_4$  এর ভর,  $w = ?$

প্রকণ-A:  $\text{H}_2\text{SO}_4$  এর মোল সংখ্যা  $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$   
 $= 0.125 \text{ mol}$

প্রকণ-B:  $\text{Na}_2\text{CO}_3$  এর মোল সংখ্যা  $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$   
 $= 0.077 \text{ mol}$

সম্প্রতিক বিক্রিয়া:  $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$

সমীকরণ হতে পাই, 1 mol  $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol  $\text{Na}_2\text{CO}_3$  বিক্রিয়া করে  $\text{H}_2\text{SO}_4$  এর 1 mol এর সাথে  
 $0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol}$

অবশিষ্ট  $\text{H}_2\text{SO}_4$  এর পরিমাণ  $= (0.125 - 0.077) \text{ mol}$   
 $= 0.048 \text{ mol}$

পাত্র-A এর প্রকণের সাথে পাত্র-B এর প্রকণের বিক্রিয়ার ফলে 0.048 mol  $\text{H}_2\text{SO}_4$  অবশিষ্ট থাকায় মিশ্রিত প্রকণের প্রকৃতি হবে অক্সিজেন।

প্রশ্ন ২ একদল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর প্রকণ প্রকৃতি এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2%  $\frac{w}{v}$  HCl এর প্রকণ প্রকৃতি নির্ণয় করে।

সি. কে. ১০।এস. ১।

- নোড কাকে বলে?
- $\text{Cl}_2$  অপেক্ষার— ব্যাখ্যা করো।
- উদাহরণের এলিটটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদাহরণের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা এক হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন গঠিত হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপেক্ষার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অণুতে তড়িৎ ঋণাত্মকতার পার্থক্য  $(3 - 3) = 0$ , তাই  $\text{Cl}_2$  বিশুদ্ধ সমযোজী ও অপেক্ষার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এহুণ ঘটে থাকে।

উদাহরণে বর্ণিত এলিটটি হলো 2%  $\left(\frac{w}{v}\right)$  HCl।

2%  $\left(\frac{w}{v}\right)$  HCl এর অর্থ হলো—

100 mL প্রকণে HCl এর ভর 2g

∴ 1000 mL " " " "  $\left(\frac{2 \times 1000}{100}\right) \text{ g}$   
 $= 20 \text{ g}$

# Answer Of Chemistry Mcq Hsc 2015

**RJ Shavelson**



## **Answer Of Chemistry Mcq Hsc 2015:**

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