Solution

1. Answer (A, B, D)

n. factor of KMnO₄ in acidic medium = 5

In faintly basic/neutral medium = 3

2. Answer (A, D)

n.f. for $C_2O_4^{-2}$ or $HC_2O_4^{-1}$ with $K_2Cr_2O_7 = 2$

n.f. for $HC_2O_4^{-1}$ with HCl = 1 and $C_2O_4^{-2}$ with HCl = 2.

Answer (C)

$$\begin{array}{l} \mathsf{MnO}_2 + (\mathsf{NH}_4)_2 \, \mathsf{SO}_4 + 2\mathsf{H}_2 \mathsf{SO}_4 \to \mathsf{MnSO}_4 + (\mathsf{NH}_4)_2 \, \mathsf{S}_2 \mathsf{O}_8 + 2\mathsf{H}_2 \mathsf{O} \\ \mathsf{n} = 2 \\ \mathsf{n} = \frac{1}{2} \end{array}$$

4. Answer (C)

Let the two metal carbonates be MCO₂ and MCO₃. Let M be 13.6% by weight

Weight of sample = 2.5 g

Weight of metal
$$M = \frac{13.6}{100} \times 2.5 = 0.34g$$

Weight of CO₂ evolved = 1.33g

Moles of CO₂ evolved =
$$\frac{1.33}{44}$$
 = 0.0302

... Weight of CO₃²⁻ in the sample = 0.0302 × 60 = 1.812 g

% of M' =
$$\frac{2.5 - 0.34 - 1.812}{2.5} \times 100 = 13.92\%$$

Answer (A. B. C).

$$\Rightarrow$$
 $n_{H_2S} \times 2 = 2.40 \times 0.02$

$$\Rightarrow$$
 n_{H-S} = 2.40 × 10⁻² m moles

... % of S in steel sample

$$= \frac{2.4 \times 10^{-3} \times 32}{0.3 \times 100} \times 100$$
$$= 0.256\%$$

6. Answer (A, B, D)

For equimolar mixture of H2SO4 and SO3

$$x + 80 + x \times 98 = 10$$

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