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web calorimeter the initial temperature of the water was 22.5 °C and the temperature of the solution after mixing was 20.4 °C on the basis of this experiment what is the heat of solution per mole of KNO_3 ? c_p of water is 4.184 J/g·K
 soln 2.58 g 98.57 g 101.15 g ΔT 20.4 °C 22.5 °C 2.1 °C

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web in a calorimeter the white solid BaSO_4 forms and the temperature of the mixture increases to 28.1 °C assuming that the calorimeter absorbs only a negligible quantity of heat and that the specific heat capacity of the solution is 4.18 J/g·°C and that the density of the final solution is 1.0 g/mL calculate the enthalpy change per mole

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web problem 8.5.1 6.8.5.1.6 when 50.0 g of 0.200 M NaCl aq at 24.1 °C is added to 100.0 g of 0.100 M AgNO_3 aq at 24.1 °C in a calorimeter the temperature increases to 25.2 °C as AgCl s forms assuming the specific heat of the solution and products is 4.20 J/g·°C calculate the approximate amount of heat in joules produced answer