

112. The IUPAC name of



is

- (1) 3, 1-diethyl-4,2,2-dimethylpentane
(2) 4, 4-dimethyl-5, 5-diethylpentane
(3) 5, 5-diethyl-4, 4-dimethylpentane
(4) 3-ethyl-4, 4-dimethylheptane

Ans.
Sol.



The correct answer is 3-ethyl-4, 4-dimethylheptane.

Hence, (4) is correct.

113. Which of the following species exhibits the diamagnetic behaviour?

- (1) O_2^+
(2) O_2
(3) O_2^-
(4) NO

Ans.
Sol.

The correct option is O_2^+

This species has 18 σ , which are filled in such a way that all molecular orbitals are fully filled, so diamagnetic.



Hence, (1) is correct.

114. The stability of divalides of Sn, Ga, In and Pb increases steadily in the sequence

- (1) $\text{GeX}_2 < \text{SnX}_2 < \text{SnX}_2 < \text{PbX}_2$
(2) $\text{SnX}_2 < \text{GaX}_2 < \text{PbX}_2 < \text{SnX}_2$
(3) $\text{SnX}_2 < \text{GaX}_2 < \text{SnX}_2 < \text{PbX}_2$
(4) $\text{PbX}_2 < \text{SnX}_2 < \text{GaX}_2 < \text{SnX}_2$

Ans.
Sol.

Due to inert pair effect, the stability of +2 oxidation state increases as we move down this group.
 $\therefore \text{SnX}_2 < \text{GaX}_2 < \text{SnX}_2 < \text{PbX}_2$

Hence, (3) is correct.

115. Identify the incorrect statement among the following

- (1) Ozone reacts with SO_2 to give SO_3
(2) Silicon reacts with $\text{NaOH}_{\text{conc}}$ in the presence of air to give Na_2SiO_3 and H_2O
(3) Cl_2 reacts with excess of NH_3 to give N_2 and HCl
(4) Br_2 reacts with hot and strong NaOH solution to give NaBr , NaBrO_3 and H_2O

Ans.
Sol.

(4) Br_2 reacts with hot and strong NaOH to give NaBr , NaBrO_2 and H_2O .

Hence, (4) is incorrect statement.

116. The charge/radius ratio of a cation determines its polarizing power. Which one of the following sequences represents the increasing order of the polarizing power of the cationic species, K^+ , Ca^{2+} , Mg^{2+} , Be^{2+} ?

- (1) $\text{Mg}^{2+}, \text{Be}^{2+}, \text{K}^+, \text{Ca}^{2+}$
(2) $\text{Be}^{2+}, \text{K}^+, \text{Ca}^{2+}, \text{Mg}^{2+}$
(3) $\text{K}^+, \text{Ca}^{2+}, \text{Mg}^{2+}, \text{Be}^{2+}$
(4) $\text{Ca}^{2+}, \text{Mg}^{2+}, \text{Be}^{2+}, \text{K}^+$

Ans.
Sol.

(3) Higher the charge/radius ratio, more is the polarizing power.



Hence, (3) is correct.

117. The density (in g mL^{-1}) of a 3.65 M sulphuric acid solution that is 20% H_2SO_4 (Molar mass = 98 g mol^{-1}) by mass will be

- (1) 1.64
(2) 1.88
(3) 1.22
(4) 1.45

Ans.
Sol.

Let the density of solution be d

Molarity of solution given = 3.6

i.e. 1 litre of solution contains 3.6 moles of H_2SO_4

or 1 litre of solution contains 3.6×98 gms of H_2SO_4

Since, the solution is 20% by mass,

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PART II Based on Class XII NCERT The Solid State Solutions Electrochemistry Chemical Kinetics Surface Chemistry Nuclear Chemistry p Block Elements The d and f Block Elements Coordination Compounds Haloalkanes and Haloarenes Alcohols Phenols and Ethers Aldehydes Ketones and Carboxylic Acids Nitrogen Containing Compounds Biomolecules Polymers Chemistry in Everyday Life Analytical Chemistry General Principles and Processes of Isolation of Elements

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 - Determining Your Reading Goals
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 - User-Friendly Interface
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 - Aieee 2014 Paper 2 Solutions Code K Enhanced eBook Features
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 - Highlighting and Note-Taking Aieee 2014 Paper 2 Solutions Code K
 - Interactive Elements Aieee 2014 Paper 2 Solutions Code K
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Aieee 2014 Paper 2 Solutions Code K
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 - Creating a Diverse Reading Collection Aieee 2014 Paper 2 Solutions Code K
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 - Minimizing Distractions
 - Managing Screen Time
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 - Carving Out Dedicated Reading Time
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