



Colligative Properties

Answer Key

Vocabulary: boiling point, colligative property, concentration, dissociate, freezing point, manometer, osmosis, osmotic pressure, solute, solution, solvent, vapor pressure

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

[Note: The purpose of these questions is to activate prior knowledge and get students thinking. Students are not expected to know the answers to the Prior Knowledge Questions.]

1. When the forecast calls for freezing rain, salt is often spread on the roads. Why is this done?

Salt is spread on the roads to prevent ice and snow from accumulating.

2. Antifreeze is added to car radiators to prevent the water in the radiators from both freezing and boiling. How do you think antifreeze affects the temperatures at which water freezes (the **freezing point**) and boils (the **boiling point**)?

Antifreeze lowers the freezing point and raises the boiling point of water.

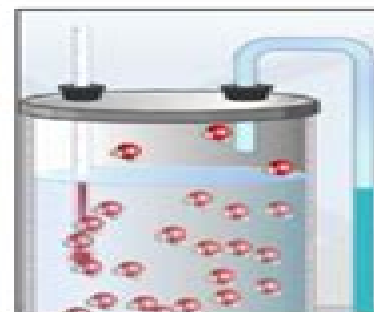
Gizmo Warm-up

Adding salt or other substances to water can affect the temperature at which it freezes or boils. These effects and others, known as **colligative properties**, are explored in the Colligative Properties Gizmo.

Check that **No solute** is selected and the **Air temp.** is 25 °C.

1. Describe the motion of the water molecules in the beaker.

The water molecules move randomly in the container. When water molecules collide, they bounce off of one another like rubber balls.



2. Change the **Air temp.** to -10 °C. What do you see as the water freezes?

As molecules freeze from the surface down, they stop moving freely and begin to vibrate in place.

3. Now change the **Air temp.** to 110 °C. What do you see as the water boils?

The molecules move more rapidly and bubbles appear in the beaker. The water disappears from the U-shaped tube that is attached to the beaker.



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