9. cobalt (III) ion (Co-3) - 27 electrons as an atom; 24 electrons as an ion.\* chromium (VI) ion (Cr<sup>+6</sup>) - 24 electrons as an atom; ✓8 electrons as an ion. 76 TH THE THE THE THE THE THE THE 2 Valence electrons (4,) and 4" a electrons or removal
11. arsenide ion (As 1) - 33 electrons as an atom; 36 electrons as an ion electrons as an ion. 12. argon (Ar) - 18 electrons 26 76 76 76 76 76 76 76 767 6 13. Use the electron orbital diagrams for questions #2 through #12 to determine how many valence electrons are contained in the elements in (a) through (f). a. #2 Ne 8 b.#3P 5 c.#4Fe 2 e. #11 As3 8 f. #12 Ar 8 An electron orbital diagram represents each individual electron of an atom or an ion with an arrow to represent electron spin. An electron configuration is not as specific of a model. An electron configuration represents the sublevels that contain electrons and the number of electrons in each sublevel is represented with a superscript (Xingerson). 14. beryllium (Be) - 4 electrons\* 1s 2, 2s 2, 2p 3s 3p 4s 3d 4p 5s 4d 5p 15. chlorine (Cl) - 17 electrons 1s 2 2s 2 2p 6 3s 2 3p 5 4s 3d 4p 5s 4d 5p 16. tellurium (Te) - 52 electrons\* 1s 2 2s2 2p6 3s2 3p6 4s2 3d/0 4p6 5s2 4d/0 5p4 Valence Electrons Core Flee

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**L** Reisser

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