

Full Name _____ Date _____ Per _____ Chemistry Chapter 6 The Periodic Table

Periodic Trends Practice

1. Identify each element as a metal, metalloid, or nonmetal

germanium _____metalloid_____

zinc _____metal_____

phosphorous _____nonmetal_____

lithium _____metal_____

2. Give two examples of elements for each category

noble gases _____He, Ne, Ar, Kr, Xe, Rn_____

halogens _____F, Cl, Br, I, At_____

alkali metals _____Li, Na, K, Rb, Cs, Fr_____

alkaline earth metals _____Be, Mg, Ca, Sr, Ba, Ra_____

transition metals _____Any elements from 3B-2B_____

3. What trend in atomic radius do you see as you go down a group/family on the periodic table? What causes this trend? Increases going down due to added energy levels. The inside energy levels are filled up so all remaining electrons

must go into open spaces further away from the nucleus.

4. What trend in atomic radius do you see as you go across a period/row on the periodic table? Decreases going across

due to electrons being added into the same valence shell. Protons would be added to the nucleus, making the positive charge stronger and electrons placed in the same space would be pulled in closer by the nucleus.

5. Circle the atom in each pair that has the largest atomic radius.

A. A B

B. S O

C. Br Cl

D. Na Al

E. O F

F. Mg Ca

6. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

O < C < Al < K

7. Define ionization energy. The amount of energy required to remove an electron from a neutral atom

8. Is it easier to form a positive ion with an element that has a high ionization energy or an element that has a low ionization energy? Explain. It would be easier to form a + ion with a low ionization energy because less energy would need to be used to remove the electron.

9. What trend in ionization energy do you see as you go down a group/family on the periodic table? Going down a group the ionization energy decreases. The valence electrons are further away and easier to take away.

10. What trend in ionization energy do you see as you go across a period/row on the periodic table? What causes this trend? Going across the period, ionization energy increases. The atoms are smaller, so the nucleus is pulling electrons in closer making them harder to remove. The atoms are also increasing their number of valence electrons. As they get closer to having a complete set of 8 valence electrons they do not want to lose electrons any more, they want to gain them instead.

11. Circle the atom in each pair that has the greater ionization energy

A. Li Be

B. Ca Ba

C. Na K

D. P Ar

E. Cl Si

F. I K

12. Define electronegativity. The ability of an atom to attract electrons in a bond.

13. What trend in electronegativity do you see as you go down a group/family on the periodic table? Values decrease within a group. Since the valence shell becomes further away, it is harder for the nucleus to draw electrons towards it.

14. What trend in electronegativity do you see as you go across a period/row on the periodic table? What causes this trend? Values increase across a period because the atoms are getting smaller and the nucleus can pull electrons in easier and the atoms are closer to being like noble gases. They want to be able to complete their valence shell.

Answer Key For Periodic Trends Review Test

**Morris Hein, Scott Pattison, Susan
Arena, Leo R. Best**



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Table of Contents Answer Key For Periodic Trends Review Test

1. Understanding the eBook Answer Key For Periodic Trends Review Test
 - The Rise of Digital Reading Answer Key For Periodic Trends Review Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key For Periodic Trends Review Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key For Periodic Trends Review Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key For Periodic Trends Review Test

- Personalized Recommendations
 - Answer Key For Periodic Trends Review Test User Reviews and Ratings
 - Answer Key For Periodic Trends Review Test and Bestseller Lists
5. Accessing Answer Key For Periodic Trends Review Test Free and Paid eBooks
 - Answer Key For Periodic Trends Review Test Public Domain eBooks
 - Answer Key For Periodic Trends Review Test eBook Subscription Services
 - Answer Key For Periodic Trends Review Test Budget-Friendly Options
 6. Navigating Answer Key For Periodic Trends Review Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key For Periodic Trends Review Test Compatibility with Devices
 - Answer Key For Periodic Trends Review Test Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key For Periodic Trends Review Test
 - Highlighting and Note-Taking Answer Key For Periodic Trends Review Test
 - Interactive Elements Answer Key For Periodic Trends Review Test
 8. Staying Engaged with Answer Key For Periodic Trends Review Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key For Periodic Trends Review Test
 9. Balancing eBooks and Physical Books Answer Key For Periodic Trends Review Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key For Periodic Trends Review Test
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Answer Key For Periodic Trends Review Test
 - Setting Reading Goals Answer Key For Periodic Trends Review Test
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Answer Key For Periodic Trends Review Test

- Fact-Checking eBook Content of Answer Key For Periodic Trends Review Test
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
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