

Name: _____ Date: _____ Period: _____

What happened to
Avogadro when he
got bit by 6.02×10^{23} mosquitoes?
He got Mol-aria

POGIL- Stoichiometry

How do chemists use balanced chemical equations?

Why?

Chemists use balanced chemical equations as a basis to calculate how much reactant is needed or product is formed in a reaction. This is called **Stoichiometry**- (stoi-key-ah-meh-tree) Another way of looking at it is using the mole ratio from the balanced equation and information about one compound in the reaction to determine information about another compound in the equation. A mole ratio is a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles. In chemical calculations, mole ratios are used to convert between moles one thing and moles of another.

Purpose: In this activity we will address the question: *How do I convert between different chemical species in a given reaction?*

MODEL 1: Mole-to-Mole Stoichiometry (1-step)

Example: mole-to-mole

Question: Given the following equation:



How many moles of H_2 are needed to produce 34.8 moles of NH_3 ?

Answer:

$$\frac{34.8 \text{ mol NH}_3}{1} \times \frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} = 52.2 \text{ mol H}_2$$

Mole ratio

OR

$$34.8 \text{ mol NH}_3 \times \frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} = 52.2 \text{ mol H}_2$$

Mole ratio

Key Questions:

1. What was the given in the example above?
2. Where did the $\frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3}$ come from?
3. Double check to see if 52.2 moles of H_2 is the correct answer. SHOW the math below.

Answer For Basic Stoichiometry Pogil Activity

David Callaghan

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Chemistry Education in the ICT Age Minu Gupta Bhowon, Sabina Jhaumeer-Laulloo, Henri Li Kam Wah, Ponnadurai Ramasami, 2009-07-21 The 20 International Conference on Chemical Education 20 ICCE which had Chemistry in the ICT Age as the theme was held from 3 to 8 August 2008 at Le Meridien Hotel Pointe aux Piments in Mauritius With more than 200 participants from 40 countries the conference featured 140 oral and 50 poster presentations The Participants of the 20 ICCE were invited to submit full papers and the latter were subjected to peer review The selected accepted papers are collected in this book of proceedings This book of proceedings encloses 39 presentations covering topics ranging from fundamental to applied chemistry such as Arts and Chemistry Education Biochemistry and Biotechnology Chemical Education for Development Chemistry at Secondary Level Chemistry at Tertiary Level Chemistry Teacher Education Chemistry and Society Chemistry Olympiad Context Oriented Chemistry ICT and Chemistry Education Green Chemistry Micro Scale Chemistry Modern Technologies in Chemistry Education Network for Chemistry and Chemical Engineering Education Public Understanding of Chemistry Research in Chemistry Education and Science Education at Elementary Level We would like to thank those who submitted the full papers and the reviewers for their timely help in assessing the papers for publication We would also like to pay a special tribute to all the sponsors of the 20 ICCE and in particular the Tertiary Education Commission <http://tec.intnet.mu> and the Organisation for the Prohibition of Chemical Weapons <http://www.opcw.org> for kindly agreeing to fund the publication of these proceedings

Process Oriented Guided Inquiry Learning (POGIL) Richard Samuel Moog, 2008 POGIL is a student centered group learning pedagogy based on current learning theory This volume describes POGIL's theoretical basis its implementations in diverse environments and evaluation of student outcomes

Calculating in Chemistry Cosmas C Mbuzi, 2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities Students in higher courses may find it a helpful revision and enhance their clarity The book completely discusses the key topics in basic stoichiometry including the mole concept reacting quantities and empirical and molecular formulas It begins with the basic concepts and formulas required to convert various quantities to moles or amount of substance This is particularly useful to a beginner Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems The chapter also describes and applies the concept of a limiting reagent Two methods of determining a limiting reagent are explained and illustrated The concept of limiting reagent is extended to reacting volumes of gases with a short cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various

circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry **A Stoichiometry Unit** David Callaghan, 2004

Formulation and Stoichiometry Emil J. Margolis, 1968 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction

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Table of Contents Answer For Basic Stoichiometry Pogil Activity

1. Understanding the eBook Answer For Basic Stoichiometry Pogil Activity
 - The Rise of Digital Reading Answer For Basic Stoichiometry Pogil Activity
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer For Basic Stoichiometry Pogil Activity
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer For Basic Stoichiometry Pogil Activity
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer For Basic Stoichiometry Pogil Activity
 - Personalized Recommendations
 - Answer For Basic Stoichiometry Pogil Activity User Reviews and Ratings

- Answer For Basic Stoichiometry Pogil Activity and Bestseller Lists
- 5. Accessing Answer For Basic Stoichiometry Pogil Activity Free and Paid eBooks
 - Answer For Basic Stoichiometry Pogil Activity Public Domain eBooks
 - Answer For Basic Stoichiometry Pogil Activity eBook Subscription Services
 - Answer For Basic Stoichiometry Pogil Activity Budget-Friendly Options
- 6. Navigating Answer For Basic Stoichiometry Pogil Activity eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer For Basic Stoichiometry Pogil Activity Compatibility with Devices
 - Answer For Basic Stoichiometry Pogil Activity Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer For Basic Stoichiometry Pogil Activity
 - Highlighting and Note-Taking Answer For Basic Stoichiometry Pogil Activity
 - Interactive Elements Answer For Basic Stoichiometry Pogil Activity
- 8. Staying Engaged with Answer For Basic Stoichiometry Pogil Activity
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer For Basic Stoichiometry Pogil Activity
- 9. Balancing eBooks and Physical Books Answer For Basic Stoichiometry Pogil Activity
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer For Basic Stoichiometry Pogil Activity
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer For Basic Stoichiometry Pogil Activity
 - Setting Reading Goals Answer For Basic Stoichiometry Pogil Activity
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer For Basic Stoichiometry Pogil Activity
 - Fact-Checking eBook Content of Answer For Basic Stoichiometry Pogil Activity
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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