- 16. Increasing the temperature generally____
 - decreases the solubility of a solid in a liquid but increases the solubility of a gas in a liquid.
 - increases the solubility of a solid in a liquid but decreases the solubility of a gas in a liquid.
 - Increases the solubility of both a solid and a gas in a liquid.
 - D. decreases the solubility of both a solid and a gas in a liquid.
- A white precipitate was formed when BaCl_{2(aq)} was added to an aqueous solution of a salt X. The precipitate dissolved in dilute HCl with rapid efferyescence. Salt X is
 - likely to contain

 A. SO₄²⁻ ions.
 - B. SO₃²- ions.
 - C. NO₃ ions.
 - D. CO₃²⁻ ions.
- Before a reaction could take place, there should be
 - A. ionization of reactant particles.
 - B. breakage of bonds of products.
 - C. breakage of bonds of reactants.
 - D. ionization of product particles.
- 19. Consider the following reaction equation:

Silicon(IV) exide is acting as

- A. a basic oxide.
- B. a reducing agent.
- an acidic oxide.
- D. an oxidizing agent.

- 20. Which of the following acids would form normal salt only?
 - A. Tetraoxosulphate(VI) acid
 - B. Tricoconitrate(V) acid
 - C. Tetraoxophosphate(V) acid
 - D. Trioxosulphate(IV) acid
 - What is the partial pressure of oxygen at s.t.p. in a gaseous mixture containing 100 cm³ of oxygen and 900 cm³ of nitrogen gas?
 - A. 1.0 arm
 - B. 0.1 artor
 - C. 9.0 atm
 - D. 0.9 over
 - Graphite is used as a dry lubricant due to the presence of
 - A. mobile electrons.
 - B. free electrons.
 - C. octahedral structures.
 - D. layered structures.
- 23. Which of the following gases is alkaline?
 - A. NO.
 - B. CO₂
 - C. NH₃
 - D. N₂O
- 24. Which of the following statements about an equilibrium system is correct?
 - Forward and backward reactions occur at the same rate.
 - The concentrations of reactants must equal that of the products.
 - C. The concentrations of reactants and products can be changed by adding a catalyst.
 - D. The concentrations of reactants and products are not affected by a change in temperature.

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- 1. Understanding the eBook Answer For Chemistry Waec 2014 F
 - The Rise of Digital Reading Answer For Chemistry Waec 2014 F
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Answer For Chemistry Waec 2014 F
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - o Features to Look for in an Answer For Chemistry Waec 2014 F
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Answer For Chemistry Waec 2014 F
 - Personalized Recommendations
 - Answer For Chemistry Waec 2014 F User Reviews and Ratings
 - Answer For Chemistry Waec 2014 F and Bestseller Lists
- 5. Accessing Answer For Chemistry Waec 2014 F Free and Paid eBooks

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- Answer For Chemistry Waec 2014 F Budget-Friendly Options
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 - Answer For Chemistry Waec 2014 F Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer For Chemistry Waec 2014 F
 - Highlighting and Note-Taking Answer For Chemistry Waec 2014 F
 - Interactive Elements Answer For Chemistry Waec 2014 F
- 8. Staying Engaged with Answer For Chemistry Waec 2014 F
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - o Following Authors and Publishers Answer For Chemistry Waec 2014 F
- 9. Balancing eBooks and Physical Books Answer For Chemistry Waec 2014 F
 - Benefits of a Digital Library
 - \circ Creating a Diverse Reading Collection Answer For Chemistry Waec 2014 F
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer For Chemistry Waec 2014 F
 - Setting Reading Goals Answer For Chemistry Waec 2014 F
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer For Chemistry Waec 2014 F
 - Fact-Checking eBook Content of Answer For Chemistry Waec 2014 F
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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