

14. You are given a bottle of saturated calcium chloride solution at 25°C. Outline the *experimental procedure* you would use to determine the *solubility* of  $\text{CaCl}_2$  in grams/L. (3 marks)

- weigh empty dish  
 - add (10.0 mL) of sat.  $\text{CaCl}_2$  soln.  
 - evaporate  
 - weigh solid  $\text{CaCl}_2$  (s)  
 - Calculate sol. in  $\frac{g}{L} = g \div L$

15. A solution of  $\text{FeBr}_3$  has a  $[\text{Br}^-]$  which is 0.36 M. Find the  $[\text{Fe}^{3+}]$  in the same solution. (1 mark)

Answer 0.12 M

16. You dissolve  $3.20 \times 10^{-2}$  moles of  $\text{Fe}(\text{NO}_3)_3$  into water to make 2.00 L of solution.

a) Determine the  $[\text{Fe}^{3+}]$  (1 mark)

Answer 0.0160 M ( $1.60 \times 10^{-2} \text{ M}$ )

b) Determine the  $[\text{NO}_3^-]$  (1 mark)

Answer 0.0480 M ( $4.80 \times 10^{-2} \text{ M}$ )

17. 30.0 mL of 0.850 M  $\text{Ba}(\text{NO}_3)_2$  is added to 50.0 mL of water. Pay attention to S.D.s in this question!

a) Find the new  $[\text{Ba}(\text{NO}_3)_2]$  in the solution. (2 marks)

$$0.850 \text{ M} \times \frac{30.0 \text{ mL}}{80.0 \text{ mL}} = 0.31875 \text{ M}$$

Answer 0.319 M

b) Find the new  $[\text{NO}_3^-]$  in the solution. (1 mark)

Answer 0.638 M

18. 67.0 mL of 0.25 M  $\text{BaCl}_2$  is mixed with 25.0 mL of 0.30 M  $\text{NaCl}$ . Determine the final  $[\text{Cl}^-]$  in the new mixture. (2 marks)

$$[\text{BaCl}_2] = 0.25 \text{ M} \times \frac{67.0 \text{ mL}}{92.0 \text{ mL}} = 0.182065 \text{ M} \quad \begin{matrix} [\text{Cl}^-] = 0.36413 \text{ M} \\ [\text{Cl}^-] = 0.08152 \text{ M} \end{matrix}$$

$$[\text{NaCl}] = 0.30 \text{ M} \times \frac{25.0}{92.0} = 0.08152 \text{ M}$$

Answer 0.45 M

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