

Name _____

Solution Stoichiometry Worksheet

Solve the following solutions Stoichiometry problems:

1. How many grams of silver chromate will precipitate when 150. mL of 0.500 M silver nitrate are added to 100. mL of 0.400 M potassium chromate?



0.150 L AgNO_3	$\frac{0.500 \text{ moles AgNO}_3}{1 \text{ L}}$	$\frac{1 \text{ moles Ag}_2\text{CrO}_4}{2 \text{ moles AgNO}_3}$	$\frac{331.74 \text{ g Ag}_2\text{CrO}_4}{1 \text{ moles Ag}_2\text{CrO}_4}$	= 12.4 g Ag ₂ CrO ₄
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$0.100 \text{ L K}_2\text{CrO}_4$	$\frac{0.400 \text{ moles K}_2\text{CrO}_4}{1 \text{ L}}$	$\frac{1 \text{ moles Ag}_2\text{CrO}_4}{1 \text{ moles K}_2\text{CrO}_4}$	$\frac{331.74 \text{ g Ag}_2\text{CrO}_4}{1 \text{ moles Ag}_2\text{CrO}_4}$	= 13.3 g Ag ₂ CrO ₄
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2. How many mL of 0.280 M barium nitrate are required to precipitate as barium sulfate all the sulfate ions from 25.0 mL of 0.350 M aluminum sulfate? (93.8 mL barium nitrate)



$0.0250 \text{ L Al}_2(\text{SO}_4)_3$	$\frac{0.350 \text{ moles Al}_2(\text{SO}_4)_3}{1 \text{ L}}$	$\frac{3 \text{ moles Ba(NO}_3)_2}{1 \text{ moles Al}_2(\text{SO}_4)_3}$	$\frac{1 \text{ L}}{0.280 \text{ moles Ba(NO}_3)_2}$	= 0.0938 L Ba(NO ₃) ₂
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3. 25.0 mL of 0.350 M NaOH are added to 45.0 mL of 0.125 M copper (II) sulfate. How many grams of copper (II) hydroxide will precipitate?



0.0250 L NaOH	$\frac{0.350 \text{ moles NaOH}}{1 \text{ L NaOH}}$	$\frac{1 \text{ moles Cu(OH)}_2}{2 \text{ moles NaOH}}$	$\frac{97.57 \text{ g Cu(OH)}_2}{1 \text{ mole Cu(OH)}_2}$	= 0.427 g Cu(OH) ₂
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0.0450 L CuSO_4	$\frac{0.125 \text{ moles CuSO}_4}{1 \text{ L NaOH}}$	$\frac{1 \text{ moles Cu(OH)}_2}{1 \text{ moles CuSO}_4}$	$\frac{97.57 \text{ g Cu(OH)}_2}{1 \text{ mole Cu(OH)}_2}$	= 0.549 g Cu(OH) ₂
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4. What volume of 0.415 M silver nitrate will be required to precipitate as silver bromide all the bromide ion in 35.0 mL of 0.128 M calcium bromide?



0.0350 L CaBr_2	$\frac{0.128 \text{ moles CaBr}_2}{1 \text{ L CaBr}_2}$	$\frac{2 \text{ moles AgNO}_3}{1 \text{ moles CaBr}_2}$	$\frac{1 \text{ L AgNO}_3}{0.415 \text{ mole AgNO}_3}$	= 0.0216 L AgNO ₃
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5. What volume of 0.496 M HCl is required to neutralize 20.0 mL of 0.809 M sodium hydroxide?



0.0200 L NaOH	$\frac{0.809 \text{ mole NaOH}}{1 \text{ L NaOH}}$	$\frac{1 \text{ mole HCl}}{1 \text{ mole NaOH}}$	$\frac{1 \text{ L HCl}}{0.496 \text{ mole HCl}}$	= 0.0326 L HCl
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Answer Key For Stoichiometry Chem Worksheet 15 6

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