

## ChemQuest 39

# Concentration

Name: \_\_\_\_\_

Date: \_\_\_\_\_

Hour: \_\_\_\_\_

**Information: Molarity**

Concentration is a term that describes the amount of solute that is dissolved in a solution. Concentrated solutions contain a lot of dissolved solute, but dilute solutions contain only a little.

**Critical Thinking Questions**

- Consider the terms "concentrated" and "dilute". Are these qualitative or quantitative terms?  
*These are qualitative terms; very general and do not include the magnitude or quantity.*
- One way of quantitatively measuring solution concentration is with units of molarity, symbolized by M. You see 1.7 liters (L) of a sodium chloride and water solution. The label on the bottle reads "1.5 M NaCl". You don't know what molarity is, but you decide to find out. After evaporating the water out of the solution you discover that there are about 149 grams of salt. Using this information, which of the following formulas is/are correct for finding molarity?

$$A) \text{ Molarity} = \frac{\text{grams of solute}}{\text{moles of solute}}$$

$$149\text{g} \div 58.5\text{g/mol} = 2.547 \text{ mol}$$

$$B) \text{ Molarity} = \frac{\text{moles of solute}}{\text{liters of solution}}$$

$$1.5 = 2.547 \text{ mol} \div 1.7 \text{ L}$$

- Using the equation you discovered in question two, calculate the molarity of each of the following solutions.  
A) A solution is prepared by dissolving 24.9 g of  $\text{CaCl}_2$  in 210 mL (which 0.210 L) of solution.

$$1.07 \text{ M}$$

$$\text{Change g to mol: } 24.9\text{g} \div 111.1\text{g/mol} = 0.224 \text{ mol}$$

$$\text{Molarity} = \text{mol} \div \text{L} = 0.224 \div 0.210 = 1.07 \text{ M}$$

- A solution contains 12.9 g of  $\text{Na}_2\text{SO}_4$  in 325 mL of solution.

$$0.279 \text{ M}$$

$$\text{Change g to mol: } 12.9 \div 142.1\text{g/mol} = 0.09078 \text{ mol}$$

$$\text{Molarity} = \text{mol} \div \text{L} = 0.09078\text{mol} \div 0.325\text{L} = 0.279 \text{ M}$$

- Verify that I need 2.15 moles of  $\text{Ca}(\text{NO}_3)_2$  to make 358 mL of a 6.00 molar solution.

$$\text{mol} = (\text{Molarity})(\text{Liters}) = (6.00\text{M})(0.358\text{L}) = 2.15 \text{ mol}$$

- Verify that it takes 80.8 g of sodium chloride to make 425 mL of a 3.25 M solution.

$$\text{mol} = (\text{Molarity})(\text{Liters}) = (3.25\text{M})(0.425\text{L}) = 1.38 \text{ mol}$$

$$(1.38 \text{ mol})(58.5\text{g/mol}) = 80.7 \text{ g}$$

# Answer Key To Chemquest 39 Concentration

**DP Hallahan**



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