

Chemical equilibrium (reversible reactions)

Comprehension Passages with Activities & Diverse Assessments.

Reading Comprehension - 9

Industrial Applications

Chemical equilibrium governs industrial reversible reactions optimizing production. The Haber process ($N_2 + 3H_2 \rightleftharpoons 2NH_3$) uses high pressure (200 atm) and moderate temperature (400°C) with iron catalysts to maximize ammonia fertilizers, balancing yield (favored by low temperature) - rate (needing heat) in sulfuric acid synthesis ($2SO_2 + O_2 \rightleftharpoons 2SO_3$). High pressure and V_2O_5 catalysts favor 90% SO_3 conversion to H_2SO_4 . Equilibrium constants guide conditions: large K reactions (e.g., 90% formation) need less CO for removing products like Fe , shifts equilibrium right, boost output. These principles, rooted since the early 20th century, rely on Le Chatelier to tweak variables, ensuring efficiency & cost-effectiveness. Equilibrium's predictability drives chem engineering from fuels to pharmaceuticals, linking their global industry.

Comprehension Questions

1. What is chemical equilibrium in reversible reactions?

2. Why is equilibrium considered dynamic?

3. How is the equilibrium constant (K_c) calculated?

4. How does Le Chatelier's Principle predict shifts?

5. How does temperature affect equilibrium?

True or False

Read each statement and circle if it is True or False.

Minimum pressure reactions stop.	✓	✗
Minimum is dynamic.	✓	✗
Low concentration rates.	✓	✗
Catalysts shift equilibrium with catalysts.	✓	✗
Temperature favors exothermic reactions.	✓	✗
High pressure favors more gas moles.	✓	✗
High reactions shift to products.	✓	✗
Subs change equilibrium position.	✓	✗
Initial pressure makes K_c .	✓	✗
A high K_c means SO_3 is formed.	✓	✗

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