

1. Answer the following in the space provided. Where required, be sure to use correct scientific notation, significant digits and including proper units of measure. (12 marks)

(a) the product(s) of the reaction of $SO_3(g) + H_2O(l) \rightarrow$

(b) the product(s) of the reaction of $Ba(s) + H_2O(l) \rightarrow$

(c) the product(s) of the reaction of $Na_2CO_3(s) + HNO_3(aq) \rightarrow$

(d) the product(s) of the reaction of $Ag(s) + Pb(NO_3)_2(aq) \rightarrow$

(e) the number of moles in 9.97×10^{25} molecules of CO_2

(f) the number of H atoms in 2.02 moles of $CH_3(CH_2)_8COOH$

(g) the number of moles in a 71.92 g sample of $MgSO_4$

(h) the mass of 3.33×10^{20} molecules of chlorine

(i) percentage composition
of ammonium phosphate

chemical formula

molar mass

percentage composition

$H_2SO_4(aq)$	
$Ba(OH)_2(aq) + H_2(g)$	
$NaNO_3(aq) + H_2O(l) + CO_2(g)$	
no rxn.	
166 mol	
2.43×10^{25} H atoms	
0.5974 mol	
0.0392 g	
$(NH_4)_3PO_4$	
149.12 $\frac{g}{mol}$	
28.19% N	8.13% H
20.77% P	42.92% O

Answer the following with complete solutions, be sure to indicate all units of measure and express final answer to correct significant digits.

2. A molecular compound containing the unknown element X has the formula X_4O_{10} . If it is known that the compound is 56.36% oxygen by mass, calculate the molar mass of element X and state its identity. (2 marks)

$$56.36\% = \frac{10(16.00 \frac{g}{mol})}{M_{X_4O_{10}}} \times 100\% \quad 283.9 \frac{g}{mol} = 4(M_X) + 10(16.00 \frac{g}{mol})$$

$$M_X = 30.97 \frac{g}{mol}$$

$$M_{X_4O_{10}} = 283.9 \frac{g}{mol}$$

$\therefore X$ is phosphorus (P)

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Chandresh Agrawal, Nandini Books



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