

Unit 5 Review: Bonding and Metals

Determine which type of bond would occur between the atoms: ionic or covalent.

Atom 1	Atom 2	Type of Bond
Aluminum	Phosphorus	ionic
Sulfur	Oxygen	covalent
Lithium	Nitrogen	ionic
Carbon	Sulfur	covalent

Write the formulas for the following compounds and then determine if the bonds are ionic or covalent. Draw the dot diagram for the compound. If it is covalent, indicate it's shape.

Magnesium Oxide
 Type: ionic
 Chemical Formula: MgO
 Dot Diagram:
 $Mg \cdot \cdot \cdot O \cdot \cdot \cdot \rightarrow [Mg]^{2+} [O]^{2-}$
 Shape (if covalent):

Carbon Tetrafluoride
 Type: covalent
 Chemical Formula: CF_4
 Dot Diagram:
 $C \cdot \cdot \cdot F \cdot \cdot \cdot F \cdot \cdot \cdot F \cdot \cdot \cdot F \cdot \cdot \cdot$
 Shape (if covalent): tetrahedral

Methane
 Type: covalent
 Chemical Formula: CH_4
 Dot Diagram:
 $C \cdot \cdot \cdot H \cdot \cdot \cdot H \cdot \cdot \cdot H \cdot \cdot \cdot H \cdot \cdot \cdot$
 Shape (if covalent): tetrahedral

Carbon Monoxide
 Type: covalent
 Chemical Formula: CO
 Dot Diagram:
 $C \cdot \cdot \cdot O \cdot \cdot \cdot$
 Shape (if covalent): linear

Barium Bromide
 Type: ionic
 Chemical Formula: $BaBr_2$
 Dot Diagram:
 $Ba \cdot \cdot \cdot Br \cdot \cdot \cdot Br \cdot \cdot \cdot \rightarrow [Ba]^{2+} [Br]^{-} [Br]^{-}$
 Shape (if covalent):

Sodium Oxide
 Type: ionic
 Chemical Formula: Na_2O
 Dot Diagram:
 $Na \cdot \cdot \cdot O \cdot \cdot \cdot Na \cdot \cdot \cdot \rightarrow [Na]^{+} [O]^{2-} [Na]^{+}$
 Shape (if covalent):

Bromic Acid
 Type: covalent
 Chemical Formula: $HBrO_3$
 Dot Diagram:
 $H \cdot \cdot \cdot Br \cdot \cdot \cdot O \cdot \cdot \cdot O \cdot \cdot \cdot O \cdot \cdot \cdot$
 Shape (if covalent): pyramidal

Nitrate Ion
 Type: covalent
 Chemical Formula: NO_3^{-}
 Dot Diagram:
 $N \cdot \cdot \cdot O \cdot \cdot \cdot O \cdot \cdot \cdot$
 Shape (if covalent): trigonal planar

Aluminum Sulfide
 Type: ionic
 Chemical Formula: Al_2S_3
 Dot Diagram:
 $Al \cdot \cdot \cdot S \cdot \cdot \cdot S \cdot \cdot \cdot S \cdot \cdot \cdot S \cdot \cdot \cdot S \cdot \cdot \cdot$
 Shape (if covalent):

Hydrogen Sulfide
 Type: covalent
 Chemical Formula: H_2S
 Dot Diagram:
 $H \cdot \cdot \cdot S \cdot \cdot \cdot H \cdot \cdot \cdot$
 Shape (if covalent): bent

Use the elements in the following list to answer the questions below: Aluminum, Calcium, Copper, Iron, Lead, Magnesium, Mercury, Nickel, Potassium, Silver, Sodium, Tin, Titanium, Zinc

- Give one of the metals from the above list that is in its toxic form is not good for the body.
Mercury
- Give one of the metals from the list above that can be found free in nature.
Silver, Mercury, Copper, Iron
- Give one of the metals from the list above that cannot be found free in nature.
Sodium, Potassium, Magnesium, Calcium, Aluminum, Lead, Nickel, Tin, Titanium, Zinc
- Name one of the strongest transition metals from the list above.
Titanium, Nickel, Aluminum
- Give one metal from the list above which is found in nature as a chloride.
Sodium, Potassium, Magnesium, Calcium
- Which metal from the list above gets its name from the Latin word maximum?
Tin
- Which metal from the list above gets its name from the Latin word hydrargyrum?
Mercury

Answer the following questions using one of the transition metals that we discussed: Aluminum, Copper, Iron, Lead, Mercury, Nickel, Silver, Tin, Titanium, Zinc

Property	Transition Metal or Metals
Good Conductors (2)	$Ag, Cu, Al, Fe, or Hg$
High Density (1)	$Hg, Pb, or Ag$
Low Density (1)	$Al or Ti$
Soft (2)	$Sn, Pb, or Ag$
Magnetic (1)	$Fe or Ni$
Good Luster (2)	$Ag, Al, Ni, or Hg$

Uses in Pure Form	Transition Metal
Fuel for Space Shuttle	Al
Car Batteries	Pb
Extreme Cold Iron Masks	Hg
Bike Frames	Ti
Wiring	Cu
Inside of Printer	Zn

Group 1 and Group 2 elements have many similarities and differences. In the table below, give a few words about each of the concepts listed.

Property or Idea	Group 1 Elements	Group 2 Elements
Reactivity with water	vigorous	mild
Density	low density	more dense than Grp 1
Occurrence in Nature	never	never
Hardness	soft	harder than Grp 1

Match the location with the metal that we need to get from this area.

Location	Metal or Metals
Bismuthville Plant, Utah	Ni
Exton, England	Mg
Dead Sea, Palestine	$Na (K, Mg, Ca)$
Morog, Peru	Ni
Salt Lake City, Utah	$Na (K, Mg, Ca)$

Fill in the chart below about the four types of bonds we learned about. Depends on electronegativity

Bond Type	Particles	Picture	Force	Melting Point	Conductivity	Other Notes	Examples
Molecular (Covalent)	nonmetal	$H - H$	weak	low	no	usually gases	carbon dioxide
Ionic	metal ion nonmetal ion	$Na^+ Cl^-$	strong	very high	no (if aqueous)	usually solids	sodium chloride
Network Covalent	nonmetal pattern	$C - C - C$	strong	very high	no	highly hardest	diamond
Metallic	metal all	$Na - Na - Na$	strong	varies	yes	ductile malleable	copper

Answer Key Ionic Bonds Review And Reinforce

Michael J. Padilla



Answer Key Ionic Bonds Review And Reinforce:

Prentice Hall Science Explorer Michael J. Padilla, 2002 **Structure & Function of the Body - E-Book** Kevin T. Patton, Frank B. Bell, Terry Thompson, Pegg L. Williamson, 2024-06-25

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The Human Body in Health & Disease - Softcover6 Kevin T. Patton,Gary A. Thibodeau,2013-01-01 Rev ed of The human body in health disease Gary A Thibodeau Kevin T Patton 5th ed c2010 **Zwitterionic hydrogels and their biomedical applications: a review** Chengcheng Li,Hao Zheng,Xingguo Zhang,Zhihua Pu,Dachao Li,2024-03-04 The coexistence of anions and cations in zwitterionic hydrogels results in electrostatic interactions between the polymer chains This structure endows zwitterionic hydrogels with higher ion sensitivity and promising properties such as anti polyelectrolyte and thermosensitive effects Hydrophilic groups on the molecular backbone give zwitterionic hydrogels good biocompatibility and they effectively resist the non specific adsorption of proteins The abundant functional groups on the molecular skeleton also facilitate the chemical modification of zwitterionic hydrogels In recent years these excellent properties have made zwitterionic hydrogels broadly interesting and they have been heavily studied for medical applications A comprehensive review will help researchers have a deeper understanding of zwitterionic hydrogels and their potential applications In this review the types functional characteristics and applications in the biomedicine of zwitterionic hydrogels are summarized in detail In addition the challenges and opportunities for using zwitterionic hydrogels for biomedical applications are discussed

Chemical Interactions ,2005 *Ion Exchange Processes for Water and Environment Management* Amita Somya,Anish Khan,2025-07-15 This book presents theory principles and applications of ion exchangers for water and environment management It begins with an introduction ion exchange equilibrium kinetics of ion exchange process and fundamental properties of ion exchangers which make them appropriate in various applications The theories underlying the operation of ion exchange resins are explained as well as the production of resin products with groups adapted to specific ions or groups of ions including principles of ion exchange process different synthetic procedures of ion exchangers and characterization techniques and the role of ion exchangers with their specific characteristics This book is an invaluable tool to analytical chemists and researchers who are interested in the applications of ion exchange materials *In Situ Molecular Pathology and Co-expression Analyses* Gerard J. Nuovo,2020-09-03 Major advancements in the field of in situ molecular pathology have occurred since publication of the first edition In Situ Molecular Pathology and Co expression Analyses Second Edition continues to teach both the molecular basis for the improvements and the actual protocols This is the unique feature that separates it from the pack of other cook book type approaches The fields of in situ hybridization and immunohistochemistry have expanded rapidly where computer based analyses systems have greatly expanded the power of these methods Further knowledge of the marked improvements in the reagents themselves since the first edition can make the difference of

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