

American Chemical Society (A.C.S.) General Chemistry Exam

75 multiple choice questions, 110 minutes

Bring a scanner, calculator and pencil(s)

- I. States of Matter**
Ideal Gas Law: $PV = nRT$, $PM = dRT$, deviations from ideal (high P, low T)
Ionic Solids: sc, fcc, bcc
Phase Diagrams: Pressure versus Temperature
Intermolecular Forces: ion-dipole, hydrogen-bonding, etc.
- II. Stoichiometry and Thermochemistry**
Hess' Law & Enthalpy: "standard states" have $\Delta H = 0$
Percent Composition: %C, %H, etc.
Limiting Reactant(s)
Stoichiometry
- III. Atomic Structure and Periodicity**
Oxidation States of "Main group" elements (Li, Be, B, etc.)
Oxidation Numbers
Atomic orbital configurations & relation to groups on periodic table (ns^2np^2 = halogens)
Periodic Trends: Electronegativity, atomic radii
Nuclear chemistry: mass number, atomic number
- IV. Molecular Structure**
Modes of bonding: ionic, covalent (both polar and nonpolar)
Lewis Dot, VSEPR, polarity, bond angles, resonance structures
- V. Solutions**
Molarity: mol solute / L solution, also calculating molarity, $C_1V_1 = C_2V_2$
Net ionic equations
Percent dissociation in water (electrolyte strength)
Molality: mol solute / kg solvent
Freezing Point Depression $\Delta T = k_{fp}$
- VI. Acids, Bases and Ionic Equilibria**
Know the "inert" acid/base ions: Cl^- , NO_3^- , Na^+ , etc.
Know general reactions for $WA + SB$, $WB + SA$, $SB + SA$
Equivalence point, finding unknown concentrations
Buffer: what makes a buffer, what a buffer is, etc.
pH: calculate pH of $WB + SA$, $SB + WA$, pure WB , pure WA
- VII. Chemical Equilibria - Molecular**
General equilibrium expressions: know how to find K
Precipitation reactions
Effect of heat upon endothermic / exothermic reactions
- VIII. Kinetics**
Rate of disappearance, rate of formation, Catalysis in mechanisms
Rate law and half-life from experimental data; Energy of activation
- IX. Thermo-dynamics**
Enthalpy, entropy, free energy
Relation of free energy to equilibrium constants
- X. Electrochemistry and Redox**
E values for electrochemical cells; E values for electrolytic cells
Balancing acidic redox reactions
Best reducing agents, oxidizing agents, definitions of oxidized and reduced
- XI. Descriptive Chemistry**
Nuclear chemistry: particle(s) emitted which result in increase of atomic number
Elemental reactivity: with water, oxygen, bases, acid, etc.
Organic chemistry: know alkenes, alkanes, etc.; also nomenclature
Coordination chemistry: coordination numbers around central metal
Ionization energy periodic trends
General nomenclature
- XII. Laboratory**
Effect of improper drying on mass measurements
Methods for finding $[Cl^-]$

American Chemical Society A C S General Chemistry Exam

Sharmistha Basu-Dutt



American Chemical Society A C S General Chemistry Exam:

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inclusion in mathematics and science classrooms It offers solutions and successful strategies for teaching and learning in mathematics and science This book covers topics such as inclusive classrooms K 12 education pre service teaching and is a useful resource for educators sociologists academicians researchers and scientists [Acs General Chemistry Study Guide - Acs Exam Prep Secrets, Full-Length Practice Test, Detailed Answer Explanations: \[Includes Step-By-Step Video Tutor](#)

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practice to help you ace your exam Publisher's Note Products purchased from 3rd party sellers are not guaranteed by the publisher for quality authenticity or access to any online entities included with the product **Making Chemistry**

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Table of Contents American Chemical Society A C S General Chemistry Exam

1. Understanding the eBook American Chemical Society A C S General Chemistry Exam
 - The Rise of Digital Reading American Chemical Society A C S General Chemistry Exam
 - Advantages of eBooks Over Traditional Books
2. Identifying American Chemical Society A C S General Chemistry Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an American Chemical Society A C S General Chemistry Exam
 - User-Friendly Interface
4. Exploring eBook Recommendations from American Chemical Society A C S General Chemistry Exam
 - Personalized Recommendations
 - American Chemical Society A C S General Chemistry Exam User Reviews and Ratings
 - American Chemical Society A C S General Chemistry Exam and Bestseller Lists

5. Accessing American Chemical Society A C S General Chemistry Exam Free and Paid eBooks
 - American Chemical Society A C S General Chemistry Exam Public Domain eBooks
 - American Chemical Society A C S General Chemistry Exam eBook Subscription Services
 - American Chemical Society A C S General Chemistry Exam Budget-Friendly Options
6. Navigating American Chemical Society A C S General Chemistry Exam eBook Formats
 - ePub, PDF, MOBI, and More
 - American Chemical Society A C S General Chemistry Exam Compatibility with Devices
 - American Chemical Society A C S General Chemistry Exam Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of American Chemical Society A C S General Chemistry Exam
 - Highlighting and Note-Taking American Chemical Society A C S General Chemistry Exam
 - Interactive Elements American Chemical Society A C S General Chemistry Exam
8. Staying Engaged with American Chemical Society A C S General Chemistry Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers American Chemical Society A C S General Chemistry Exam
9. Balancing eBooks and Physical Books American Chemical Society A C S General Chemistry Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection American Chemical Society A C S General Chemistry Exam
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine American Chemical Society A C S General Chemistry Exam
 - Setting Reading Goals American Chemical Society A C S General Chemistry Exam
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of American Chemical Society A C S General Chemistry Exam
 - Fact-Checking eBook Content of American Chemical Society A C S General Chemistry Exam
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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