

অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পরীক্ষার কপূর, এখানে কোর্ড পরীক্ষা, আরও কয়েকটি পরীক্ষার কপূর-সম্পর্কিত কয়েকটি পরীক্ষা এবং মতামত টিউটর পরীক্ষার কপূর উত্তর অধ্যয়নিক সেবা করে। এগুলোর অনুশীলন করতে তুমি এ অধ্যয়নিক থেকে সেবায় সফল হতে পারবে।



পাত্র-A



পাত্র-B

- ক. মোলার প্রবণ কী? ১
 খ. "কোনও পানীয় লবণের চাকতা, তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করো। ২
 গ. A-পাত্রে বিদ্যমান এলিটের ভর কত নির্ণয় করো। ৩
 ঘ. ট্রাবল দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অম্লীয় বা ক্ষারীয় হবে— বিশ্লেষণ করো। ৪
১ নং প্রশ্নের উত্তর

১০০০ mL ট্রাবলে কোন ট্রাবলে 1 mol দ্রবীভূত থাকলে উক্ত ট্রাবলে মোলার প্রবণ কত।

কোনো পানীয় লবণের চাকতা তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, তাপমাত্রা বৃদ্ধির সাথে সাথে চাকতা বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট তাপমাত্রার কোনো পানীয় লবণ নিষ্ক্রিয় হয়। তখন চাকতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পানিতে দ্রবীভূত করার সময় বিয়োজন প্রক্রিয়াটি তাপস্বরী বিক্রিয়া কিন্তু 32.4°C তাপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা বৃদ্ধিজনিত কারণে চাকতার হ্রাস ঘটে।

আমরা জানি,
 $w = SMV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$
 অতএব, দ্রবীভূত H_2SO_4 এর ভর = 12.25 g

উদ্দীপকের A-পাত্রে অবস্থিত, H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$
 ট্রাবলের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
 ট্রাবলের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$
 দ্রবীভূত H_2SO_4 এর ভর, $w = ?$

১১. পাত্র-A: H_2SO_4 এর মোল সংখ্যা = $\left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

পাত্র-B: Na_2CO_3 এর মোল সংখ্যা = $\left(0.11 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$

সম্ভব বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$
 সমীকরণ হতে পাই, 1 mol H_2SO_4 = 1 mol Na_2CO_3
 1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 0.077 mol^* " " " " " " " " " " " "
 $(0.077) \text{ mol}^*$ " " " " " " " " " " " "
 $= 0.077 \text{ mol}^*$ " " " " " " " " " " " "

অবশিষ্ট H_2SO_4 এর পরিমাণ = $(0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

পাত্র-A এর ট্রাবলের সাথে পাত্র-B এর ট্রাবলের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত ট্রাবলের প্রকৃতি হবে অম্লীয়।

একমাত্র ঘনমাত্রা ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর ট্রাবলে পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\left(\frac{w}{v}\right)$ HCl এর দ্রবণ তৈরি করা হবে।

- ক. লোভ কাকে বলে? ১
 খ. Cl_2 অপোলার— ব্যাখ্যা করো। ২
 গ. উদ্দীপকের এলিটটির মোলার ঘনমাত্রা নির্ণয় করো। ৩
 ঘ. উদ্দীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো। ৪
২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেকট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে লোভ বলে।

সমযোজী বন্ধনে আকর্ষণ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা কম হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন বিশুদ্ধ ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অপূর্ণ তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_2 বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এরূপ থাকে।

উদ্দীপকে বর্ণিত এলিটটি হলো 2% $\left(\frac{w}{v}\right)$ HCl।

2% $\left(\frac{w}{v}\right)$ HCl এর অর্থ হলো—

100 mL ট্রাবলে HCl এর ভর 2g
 $\therefore 1000 \text{ mL} \dots \dots \dots \left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

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Wolfgang Guggemos



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