

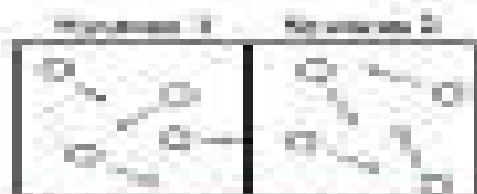
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v is $\frac{1}{2}mv^2$, so the more of particles that move fast enough, the more kinetic energy we will see for a system. Therefore, the temperature is directly proportional to the average kinetic energy of the particles in a system. The Boltzmann-probability distribution curves show that the average kinetic energy will be the same for all three systems (the area under the curves is the same).

3. Sketch a curve to represent the relative distribution of the particles in the following two gases, samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are more spread out in the box, so all the individual particles have less kinetic energy. According to the equation $KE = \frac{1}{2}mv^2$, all of the particles have the same mass, so the more spread out they are, the more KE they have. In System 2, the particles are more spread out, so they have more KE. Therefore, System 1 is at a lower temperature than System 2.
- b. In which direction will kinetic energy be transferred? Justify your answer in terms of molecular collisions.
KE will be transferred from System 1 to System 2. The particles in System 1 have less KE than the particles in System 2. When the particles collide, the particles in System 1 will transfer some of their KE to the particles in System 2. This will happen until the particles in System 1 have the same KE as the particles in System 2.

Answer Key Unit 6 Thermochemistry

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