

Question Number	Acceptable Answers	Reject	Mark																									
23(a)	<table><tr><th>Element</th><th>%</th><th></th><th>mol</th><th>Ratio</th></tr><tr><td>C</td><td>40.44</td><td>÷ 12</td><td>= 3.37</td><td>2.999</td></tr><tr><td>H</td><td>7.87</td><td>÷ 1</td><td>= 7.87</td><td>7.004</td></tr><tr><td>O</td><td>35.96</td><td>÷ 16</td><td>= 2.2475</td><td>2.000</td></tr><tr><td>N</td><td>15.73</td><td>÷ 14</td><td>= 1.12357</td><td>1.000</td></tr></table> (1) (1) Empirical formula = C₃H₇O₂N (1) Symbols in any order Stand alone mark No TE on incorrect ratio I IGNORE significant figure and rounding errors except 1sf in mole calculation	Element	%		mol	Ratio	C	40.44	÷ 12	= 3.37	2.999	H	7.87	÷ 1	= 7.87	7.004	O	35.96	÷ 16	= 2.2475	2.000	N	15.73	÷ 14	= 1.12357	1.000		3
Element	%		mol	Ratio																								
C	40.44	÷ 12	= 3.37	2.999																								
H	7.87	÷ 1	= 7.87	7.004																								
O	35.96	÷ 16	= 2.2475	2.000																								
N	15.73	÷ 14	= 1.12357	1.000																								

Question Number	Acceptable Answers	Reject	Mark
23(b)(i)	Peak at $m/e = 89$ labelled M^+ ALLOW Any clear label e.g. $C_3H_7O_2N^+$		1

Question Number	Acceptable Answers	Reject	Mark
23(b)(ii)	$M_r = 89 = M_r(C_3H_7O_2N)$ so molecular formula is $C_3H_7O_2N$ Symbols in any order IGNORE structural and displayed formulae	Answer with no explanation $M_r = 90$ $C_3H_7O_2N^+$	1

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Albert A Gayle



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Table of Contents Answer Sheet Of 2014 C12 Ial Paper

1. Understanding the eBook Answer Sheet Of 2014 C12 Ial Paper
 - The Rise of Digital Reading Answer Sheet Of 2014 C12 Ial Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Sheet Of 2014 C12 Ial Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Sheet Of 2014 C12 Ial Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Sheet Of 2014 C12 Ial Paper
 - Personalized Recommendations
 - Answer Sheet Of 2014 C12 Ial Paper User Reviews and Ratings
 - Answer Sheet Of 2014 C12 Ial Paper and Bestseller Lists
5. Accessing Answer Sheet Of 2014 C12 Ial Paper Free and Paid eBooks
 - Answer Sheet Of 2014 C12 Ial Paper Public Domain eBooks
 - Answer Sheet Of 2014 C12 Ial Paper eBook Subscription Services
 - Answer Sheet Of 2014 C12 Ial Paper Budget-Friendly Options
6. Navigating Answer Sheet Of 2014 C12 Ial Paper eBook Formats

- ePub, PDF, MOBI, and More
 - Answer Sheet Of 2014 C12 Ial Paper Compatibility with Devices
 - Answer Sheet Of 2014 C12 Ial Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Sheet Of 2014 C12 Ial Paper
 - Highlighting and Note-Taking Answer Sheet Of 2014 C12 Ial Paper
 - Interactive Elements Answer Sheet Of 2014 C12 Ial Paper
 8. Staying Engaged with Answer Sheet Of 2014 C12 Ial Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Sheet Of 2014 C12 Ial Paper
 9. Balancing eBooks and Physical Books Answer Sheet Of 2014 C12 Ial Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Sheet Of 2014 C12 Ial Paper
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Answer Sheet Of 2014 C12 Ial Paper
 - Setting Reading Goals Answer Sheet Of 2014 C12 Ial Paper
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Answer Sheet Of 2014 C12 Ial Paper
 - Fact-Checking eBook Content of Answer Sheet Of 2014 C12 Ial Paper
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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