Chapter 1 - Chemical Foundations

 Consider the element indium, atomic number 49, atomic mass 114.8 g. The nucleus of an atom of indium-112 contains:

- a. 49 protons, 63 neutrons, 49 electrons.
- b. 49 protons, 49 neutrons.
- c. 49 protons, 49 alpha particles.
- d. 49 protons, 63 neutrons.
- e. 49 protons, 112 neutrons.

ANSWER:	d
POINTS:	1

2. A hypothetical element consists of two isotopes of masses 78.95 amu and 80.95 amu with abundances of 32.9% and 67.1%, respectively. What is the average atomic mass of this element?

a.	79:95 amu
b.	78.95 amu
C.	79.6 amu
d.	80.3 amu
e.	80.95 amu

ANSWER:	d
POINTS:	1

Naturally occurring element X exists in three isotopic forms: X-28 (27.977 amu, 92.23% abundance), X-29 (28.976 amu, 4.67% abundance), and X-30 (29.974 amu, 3.10% abundance). Calculate the atomic weight of X.

a.	28.09 amu
b.	48.63 amu
c.	27.16 amu
d.	28.97 amu
40.0	86.93 amu

ANSWER:	a a	
POINTS:		ı

4. The atomic mass of rhenium is 186.2. Given that 37.1% of natural rhenium is rhenium-185, what is the other stable isotope?

a.	113 Re
b.	187 Re
C.	189 R.e
d.	181 R.e
e.	190 R.e

ANSWER:	1	6
POINTS:	Si .	

Naturally occurring element X exists in three isotopic forms: X-28 (27.977 amu, 92.23% abundance), X-29
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Answers For Chapter One Chemical Foundations Exercises

Kenneth W. Whitten

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