

Chemistry Benchmark Assessment

Student Name: _____

1. Which of the following reactions represents a synthesis reaction?

- ☐ A. $2 \text{H}_2\text{O} (\text{g}) \rightarrow 2 \text{H}_2 (\text{g}) + \text{O}_2 (\text{g})$
- ☐ B. $\text{CaCO}_3 (\text{s}) + 2 \text{HNO}_3 (\text{aq}) \rightarrow \text{Ca(NO}_3)_2 (\text{aq}) + 2 \text{H}_2\text{O} (\text{l})$
- ☐ C. $\text{Zn} (\text{s}) + \text{CuSO}_4 (\text{aq}) \rightarrow \text{ZnSO}_4 (\text{aq}) + \text{Cu} (\text{s})$
- ☐ D. $2 \text{Mg} (\text{s}) + \text{O}_2 (\text{g}) \rightarrow 2 \text{MgO} (\text{s})$

2. When a nitrogen atom covalently bonds to three hydrogen atoms to form a molecule, it will retain one pair of unshared electrons. According to VSEPR theory, what is the shape of an ammonia molecule?

- ☐ A. trigonal planar
- ☐ B. linear tetrahedral
- ☐ C. pyramidal
- ☐ D. tetrahedral

3. An unknown isotope has 17 neutrons and an atomic mass of 32. What is the name of this element?

- ☐ A. phosphorus
- ☐ B. sodium
- ☐ C. sulfur
- ☐ D. chlorine

4. Which of the following does not possess the same number of valence electrons?

- ☐ A. elements within the same period
- ☐ B. all elements with the same atomic number
- ☐ C. all elements with the same atomic mass
- ☐ D. elements within the same group

Answers For Chemistry Benchmark

David R. Klein



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