

CHAPTER 10: Reactions

Reaction Energy**SECTION 1**

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space on the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ of a certain compound is 2074°C . Absorption of 4.000 J of energy will result in the decomposition of _____.

- _____ A compound has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If its specific enthalpy of formation is measured as 1.0 kJ/mol at 25°C , it will absorb _____ J of energy when 100 g of it is heated from 25°C to 100°C .

- The product is a molecular form of oxygen, and the element has an enthalpy of 0.001 kJ/mol .

- What is the value of ΔH for this reaction?

Answers For Reaction Energy Section 1

**Christopher R. Hren, John T.
Moore, Peter J. Mikulecky**



Answers For Reaction Energy Section 1:

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