

THE CHEMISTRY OF LIFE

Living matter is composed of many different elements that are made up of small units called atoms. An atom of any element has a nucleus that is surrounded by particles called electrons. Each electron has a negative charge. A nucleus usually has two types of particles, protons and neutrons. Each proton has a positive charge. Neutrons have no charge. The number of protons in a nucleus of atom is called its atomic number. When you add up the number of protons and neutrons together, you get what is called the atomic mass. The chart below shows six elements that make up most of your body.

Use this information about atoms and the facts in the chart to answer the questions that follow the chart.

Element	Element Symbol	Atomic number	Atomic mass	Percent in the human body
Hydrogen	H	1	1	9.5
Carbon	C	6	12	18.5
Nitrogen	N	7	14	3.0
Oxygen	O	8	16	65.0
Phosphorus	P	15	31	1
Sulfur	S	16	32	0.3

1. What element is the most abundant in the human body? What is its symbol?

2. What is the atomic number of the element in question 1?

3. What makes up to 18.5 percent of the body? What is its symbol?

4. Write the atomic number of phosphorous. How many protons does it have?

5. Finish the word equation

6. Atomic mass = the number of _____ + the number of _____
7. What is the atomic mass of nitrogen?

8. How many neutrons are there in the nucleus of a nitrogen atom?

9. Write the symbol of the element found in only 0.3 percent of human body.

10. Which element in the chart has no neutrons?

Answer Key To The Chemistry Of Life

Oswaal Editorial Board



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Table of Contents Answer Key To The Chemistry Of Life

1. Understanding the eBook Answer Key To The Chemistry Of Life
 - The Rise of Digital Reading Answer Key To The Chemistry Of Life
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key To The Chemistry Of Life
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key To The Chemistry Of Life
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key To The Chemistry Of Life
 - Personalized Recommendations
 - Answer Key To The Chemistry Of Life User Reviews and Ratings

- Answer Key To The Chemistry Of Life and Bestseller Lists
- 5. Accessing Answer Key To The Chemistry Of Life Free and Paid eBooks
 - Answer Key To The Chemistry Of Life Public Domain eBooks
 - Answer Key To The Chemistry Of Life eBook Subscription Services
 - Answer Key To The Chemistry Of Life Budget-Friendly Options
- 6. Navigating Answer Key To The Chemistry Of Life eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key To The Chemistry Of Life Compatibility with Devices
 - Answer Key To The Chemistry Of Life Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key To The Chemistry Of Life
 - Highlighting and Note-Taking Answer Key To The Chemistry Of Life
 - Interactive Elements Answer Key To The Chemistry Of Life
- 8. Staying Engaged with Answer Key To The Chemistry Of Life
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key To The Chemistry Of Life
- 9. Balancing eBooks and Physical Books Answer Key To The Chemistry Of Life
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key To The Chemistry Of Life
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key To The Chemistry Of Life
 - Setting Reading Goals Answer Key To The Chemistry Of Life
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key To The Chemistry Of Life
 - Fact-Checking eBook Content of Answer Key To The Chemistry Of Life
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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