Pre-lab Questions for Experiment #5

Rates of Chemical Reactions I: A Clock Reaction

1) Complete and balance the following reactions:

(a)
$$S_2O_8^{2-} + \Gamma \rightarrow$$

(b)
$$I_2 + S_2O_3^{2-} \rightarrow$$

2) The rate law for the reaction

$$2H_2(g) + 2NO(g) \longrightarrow N_2(g) + 2H_2O(g)$$

is rate = k[H₂][NO]². Which of the following mechanisms can be ruled out on the basis of the observed rate expression?

$$\begin{array}{c} \textit{Mechanism I} \\ \text{H}_2 + \text{NO} \longrightarrow \text{H}_2\text{O} + \text{N} & (\text{slow}) \\ \text{N} + \text{NO} \longrightarrow \text{N}_2 + \text{O} & (\text{fast}) \\ \text{O} + \text{H}_2 \longrightarrow \text{H}_2\text{O} & (\text{fast}) \\ \end{array}$$

$$\begin{array}{c} \textit{Mechanism II} \\ \text{H}_2 + 2\text{NO} \longrightarrow \text{N}_2\text{O} + \text{H}_2\text{O} & (\text{slow}) \\ \text{N}_2\text{O} + \text{H}_2 \longrightarrow \text{N}_2 + \text{H}_2\text{O} & (\text{fast}) \\ \end{array}$$

$$\begin{array}{c} \textit{Mechanism III} \\ 2\text{NO} \Longrightarrow \text{N}_2\text{O}_2 & (\text{fast equilibrium}) \\ \text{N}_2\text{O}_2 + \text{H}_2 \longrightarrow \text{N}_2\text{O} + \text{H}_2\text{O} & (\text{slow}) \\ \end{array}$$

$$\begin{array}{c} \text{N}_2\text{O}_2 + \text{H}_2 \longrightarrow \text{N}_2\text{O}_2 + \text{H}_2\text{O} & (\text{fast}) \\ \end{array}$$

3) Consider the reaction

From the following data obtained at a certain temperature, determine the order of the reaction and calculate the rate constant:

[A] (M)	[B] (M)	Rate (M/s)
1.50	1.50	3.20×10^{-1}
1.50	2.50	3.20×10^{-1}
3.00	1.50	6.40×10^{-1}

4) The following gas-phase reaction was studied at 290°C by observing the change in pressure as a function of time in a constant-volume vessel:

$$CICO_2CCI_3(g) \longrightarrow 2COCI_2(g)$$

Determine the order of the reaction and the rate constant based on the following data:

Time (s)	Partial Pressure of CICO2CCI3 (mmHg)	
0	15.76	
181	12.64	
513	8.73	
1164	4.44	

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