

Question 1: Water at 20°C falls from a height of 854 m. If the whole energy is used in increasing the temperature, find out the final temperature. Specific heat of water is 4200 J/K kg .

Given that

$$T_1 = 20^{\circ}\text{C} = 20 + 273 = 293 \text{ K}$$

$$\text{height} = h = 854 \text{ m}$$

$$\text{Specific heat of water} = 4200 \text{ J K}^{-1} \text{ kg}^{-1}$$

$$\text{Final temperature} = T_2 = ?$$

$$\text{P.E} = mgh \quad \text{--- (1)}$$

Solution:

We know that

$$\Delta Q = cm \Delta T \quad \text{--- (2)}$$

Whole energy is used in increasing the temperature (Given)

$$\Delta Q = \text{P.E}$$

From (1) and (2)

$$cm \Delta T = mgh$$

$$c \Delta T = gh$$

$$c (T_2 - T_1) = gh$$

$$T_2 - T_1 = \frac{gh}{c}$$

$$T_2 = \frac{gh}{c} + T_1$$

$$= \frac{9.8 \times 854}{4200} + 293$$

$$= \frac{8369.2}{4200} + 293 = 1.99 + 293$$

$$T_2 = 294.99 \text{ K} = 295 - 273^{\circ}\text{C}$$

$$T_2 = 22^{\circ}\text{C}$$

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Jasjeet Singh Bagla, Sunu Engineer



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Table of Contents Answer The Following Questions About The Thermodynamics Of

1. Understanding the eBook Answer The Following Questions About The Thermodynamics Of
 - The Rise of Digital Reading Answer The Following Questions About The Thermodynamics Of
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer The Following Questions About The Thermodynamics Of
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer The Following Questions About The Thermodynamics Of
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer The Following Questions About The Thermodynamics Of
 - Personalized Recommendations
 - Answer The Following Questions About The Thermodynamics Of User Reviews and Ratings
 - Answer The Following Questions About The Thermodynamics Of and Bestseller Lists
5. Accessing Answer The Following Questions About The Thermodynamics Of Free and Paid eBooks

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- Answer The Following Questions About The Thermodynamics Of Public Domain eBooks
- Answer The Following Questions About The Thermodynamics Of eBook Subscription Services
- Answer The Following Questions About The Thermodynamics Of Budget-Friendly Options
- 6. Navigating Answer The Following Questions About The Thermodynamics Of eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer The Following Questions About The Thermodynamics Of Compatibility with Devices
 - Answer The Following Questions About The Thermodynamics Of Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer The Following Questions About The Thermodynamics Of
 - Highlighting and Note-Taking Answer The Following Questions About The Thermodynamics Of
 - Interactive Elements Answer The Following Questions About The Thermodynamics Of
- 8. Staying Engaged with Answer The Following Questions About The Thermodynamics Of
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer The Following Questions About The Thermodynamics Of
- 9. Balancing eBooks and Physical Books Answer The Following Questions About The Thermodynamics Of
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer The Following Questions About The Thermodynamics Of
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer The Following Questions About The Thermodynamics Of
 - Setting Reading Goals Answer The Following Questions About The Thermodynamics Of
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer The Following Questions About The Thermodynamics Of
 - Fact-Checking eBook Content of Answer The Following Questions About The Thermodynamics Of
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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