

### **Ionic and Covalent Naming Practice **Answers****

*For each of the following questions, determine whether the compound is ionic or covalent and name it appropriately.*

- |     |                           |                                  |
|-----|---------------------------|----------------------------------|
| 1)  | $\text{Na}_2\text{CO}_3$  | <b>sodium carbonate</b>          |
| 2)  | $\text{P}_4\text{O}_{10}$ | <b>tetraphosphorus decaoxide</b> |
| 3)  | $\text{N}_2\text{O}_5$    | <b>dinitrogen pentoxide</b>      |
| 4)  | $\text{FeSO}_4$           | <b>iron (II) sulfate</b>         |
| 5)  | $\text{SiO}_2$            | <b>silicon dioxide</b>           |
| 6)  | $\text{FeCl}_3$           | <b>iron (III) chloride</b>       |
| 7)  | $\text{CoBr}_2$           | <b>cobalt (II) bromide</b>       |
| 8)  | $\text{B}_2\text{H}_4$    | <b>diboron tetrahydride</b>      |
| 9)  | $\text{CO}$               | <b>carbon monoxide</b>           |
| 10) | $\text{PCl}_3$            | <b>phosphorus trichloride</b>    |

*For each of the following questions, determine whether the compound is ionic or covalent and write the appropriate formula for it.*

- |     |                        |                                                      |
|-----|------------------------|------------------------------------------------------|
| 11) | dinitrogen trioxide    | <b><math>\text{N}_2\text{O}_3</math></b>             |
| 12) | nitrogen dioxide       | <b><math>\text{NO}_2</math></b>                      |
| 13) | sulfur pentoxide       | <b><math>\text{SO}_5</math></b>                      |
| 14) | lithium acetate        | <b><math>\text{LiC}_2\text{H}_3\text{O}_2</math></b> |
| 15) | phosphorus trifluoride | <b><math>\text{PF}_3</math></b>                      |
| 16) | vanadium (V) oxide     | <b><math>\text{V}_2\text{O}_5</math></b>             |
| 17) | aluminum hydroxide     | <b><math>\text{Al}(\text{OH})_3</math></b>           |
| 18) | zinc sulfide           | <b><math>\text{ZnS}</math></b>                       |
| 19) | silicon tetrafluoride  | <b><math>\text{SiF}_4</math></b>                     |
| 20) | silver phosphate       | <b><math>\text{Ag}_3\text{PO}_4</math></b>           |

# Answers Ionic And Covalent Bond Practice Chemfiesta

**S Nieto**



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**Chemical Bonding** M.S. Sethi & M. Satake, 2010 Contents Chemical Bonding I Basic Concepts Chemical Bonding II Additional Aspects Intermolecular Force and Crystal Structures

**Covalent Bond 35 Success Secrets - 35 Most Asked Questions on Covalent Bond - What You Need to Know** Shirley Soto, 2014-10-14 A Blue Ribbon Covalent bond Guide A covalent bond is a biochemical bond that includes the parting of negatron matches amid particles The steady level of alluring and hideous drives amid particles once they share electrons is recognized like covalent joining For numerous particles the parting of electrons permits every one particle to attain the equal of a complete outside shell comparable to a steady microelectronic arrangement There has never been a Covalent bond Guide like this It contains 35 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Covalent bond A quick look inside of some of the subjects covered Chemical bonding Covalent bond Noncovalent bonding Drug Design Covalent bond History Noncovalent bonding Cation Anion Noncovalent bonding Hydrophobic effect Polar covalent bond Polarity of bonds Noncovalent bonding Boiling Points of Liquids Noncovalent bonding London Dispersion Forces Coordinate Covalent Bond Examples Noncovalent bonding effects Noncovalent bonding H bonding Noncovalent bonding Polar Polar covalent bond Polar molecules Covalent bonds Polarity of covalent bonds Polar covalent bond Polarity of molecules Noncovalent bonding Interaction Noncovalent bonding Van der Waals Forces Covalent bonds Subdivision of covalent bonds Covalent Bond Classification Polar covalent bond Hybrids Noncovalent bonding Electrostatic Interactions Polar covalent bond Nonpolar molecules Covalent Bond Classification Other uses Covalent bonds History Noncovalent bonding Dipole Dipole Noncovalent bonding Protein Folding Structure and much more

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then applied to a variety of covalent molecules with discussions of bond lengths and angles and hence molecular shapes Extensive comparisons of valence bond theory and VSEPR theory with molecular orbital theory are included Metallic bonding is related to electrical conduction and semi conduction The energetics of ionic bond formation and the transition from ionic to covalent bonding is also covered

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**The Chemical Bond** Gernot Frenking,Sason Shaik,2014-06-13 This is the perfect complement to Chemical Bonding Across the Periodic Table by the same editors who are two of the top scientists working on this topic each with extensive experience and important connections within the community The resulting book is a unique overview of the different approaches used for describing a chemical bond including molecular orbital based valence bond based ELF AIM and density functional based methods It takes into account the many developments that have taken place in the field over the past few decades due to the rapid advances in quantum chemical models and faster computers

**The Covalent Bond** Henry Sinclair Pickering,1977 Polar Covalence R Sanderson,2012-12-02 Polar Covalence provides a detailed account of a successful approach to understanding chemistry from knowledge of atomic structure and the properties that result from this structure This book discusses the nature of multiple bonds Organized into 16 chapters this book begins with an overview of the interrelationships of various basic atomic properties This text then describes chemical bonding which can only occur when the nuclei of both atoms can attract the same electrons Other chapters consider the bond energy of multiple bonds which can be determined by calculating the energy in the usual way as though the bonds were single but of the experimental length This book discusses as well the reduction of the lone pair bond weakening effect through the formation of multiple bonds The final chapter deals with the relative roles of principles and practice in the

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